



# **Safe Discharge of Landfill Leachate to the Environment**

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Report #

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# TABLE OF CONTENTS

1. INTRODUCTION .....	3
1.1 BACKGROUND .....	3
1.2 WATER QUALITY REGULATIONS.....	3
1.3 LEACHATE WATER QUALITY.....	4
1.4 POLLUTANTS OF CONCERN.....	6
1.5 LEACHATE QUANTITIES.....	7
1.6 METHODS OF LEACHATE MANAGEMENT .....	9
1.7 ADVANCED OXIDATION PROCESSES.....	13
1.8 OBJECTIVES.....	4
2. METHODOLOGY .....	5
2.1 LEACHATE COLLECTION .....	5
2.2 PILOT SCALE REACTOR.....	8
2.3 FALLING FILM REACTOR .....	10
2.4 FLOW THROUGH REACTOR.....	11
2.5 IMPROVEMENTS/ADDITIONS .....	12
2.6 EXPERIMENT PROCEDURE.....	14
2.7 TiO <sub>2</sub> DOSING.....	15
2.8 CRITICAL ORIFICE ADVANCED OXIDATION PROCESS.....	17
2.9 CRYSTAL VIOLET TEST .....	18
2.10 IMPROVING COD REMOVAL.....	19
2.11 ANALYTICAL METHODS FOR PARAMETERS OF INTEREST.....	19
2.11.1 COD.....	19
2.11.2 Ammonia-Nitrogen.....	20
2.11.3 Alkalinity .....	21

2.11.4	pH.....	21
2.11.5	Dissolved Oxygen.....	22
2.11.5	Temperature.....	22
2.12	CATALYST RECOVERY.....	22
2.13	INTENSITY OF UV LIGHT.....	23
3.	RESULTS AND DISCUSSION.....	24
3.1	FALLING FILM EXPERIMENT #1.....	24
3.2	FLOW THROUGH EXPERIMENT #2.....	29
3.3	FLOW THROUGH WITH AERATION EXPERIMENT #3.....	34
3.4	FLOW THROUGH WITH AERATION EXPERIMENT #4 WITH 0.2-g/L TiO <sub>2</sub> USING THE 150-W LAMP.....	39
3.5	TiO <sub>2</sub> DOSING TEST.....	44
3.6	CRITICAL ORIFICE ADVANCED OXIDATION PROCESS TESTING.....	45
3.7	CRYSTAL VIOLET TEST.....	46
3.8	IMPROVING COD REMOVAL.....	46
3.9	INTENSITY OF UV LIGHT.....	47
3.10	CATALYST RECOVERY.....	49
4.	CONCLUSION.....	50
4.1	SUMMARY.....	50
4.2	RECOMMENDATIONS.....	50
5.	REFERENCES.....	51

# LIST OF FIGURES

Figure 1: Dyer Park Historical Leachate Quantities (SWA personal communication 2014).	9
Figure 2. Dyer Park Landfill sample collection point.	5
Figure 3. Dyer Park sampling point pump station dry well.	6
Figure 4. Dyer Park sampling port at the pump station.	6
Figure 5. Purging the sample port prior to collection.	7
Figure 6. Filling 5-gallon sample containers.	8
Figure 7. Main components for the pilot scale reactor.	9
Figure 8. Irradiation spectrum for the Strahler NNI 125/84 XL low pressure UV lamp as provided by the manufacturer.	10
Figure 9: Reactor configured as a flow through.	11
Figure 10. Photograph of the 304 stainless steel cooling coil.	12
Figure 11. Photograph of custom flow through reactor lid.	13
Figure 12. Photograph of ball valve.	13
Figure 13. Photograph of aeration system.	14
Figure 14. Measuring leachate in a 2000-mL graduated cylinder.	15
Figure 15: Dilutions created for a standard Beer's Law curve.	17
Figure 16: Full spectrum absorption curve from 200-400-nm using the DR4000U.	17
Figure 17. Alternative AOP device.	18
Figure 18. Improved COD removal test chamber setup.	19
Figure 19. Reagent addition to cuvette for ammonia testing.	20
Figure 20. Sample reading from the HI 95715 Ammonia Medium Range ISM.	21
Figure 21. Sedimentation time lapse series over a 10-minute period.	22
Figure 22. Falling Film Reactor Temperature Comparison.	26
Figure 23. Falling Film Reactor pH Comparison.	27
Figure 24. Falling Film Reactor Ammonia Comparison.	27
Figure 25. Falling Film Reactor Alkalinity Comparison.	28
Figure 26. Falling Film Reactor COD Comparison.	28
Figure 27. Flow Through Reactor Temperature Comparison.	31
Figure 28. Flow Through Reactor pH Comparison.	32
Figure 29. Flow Through Reactor Ammonia Comparison.	32
Figure 30. Flow Through Reactor Alkalinity Comparison.	33
Figure 31. Flow Through Reactor COD Comparison.	34

Figure 32. Flow Through Reactor with Aeration Temperature Comparison.....	36
Figure 33. Flow Through Reactor with Aeration pH Comparison. ....	37
Figure 34. Flow Through Reactor with Aeration Ammonia Comparison. ....	37
Figure 35. Flow Through Reactor with Aeration Alkalinity Comparison. ....	38
Figure 36. Flow Through Reactor with Aeration COD Comparison.....	38
Figure 37. Flow Through Reactor with 150-W Lamp and Aeration using $\text{TiO}_2 = 0.2 \text{ g/L}$ Temperature Trend.....	41
Figure 38. Flow Through Reactor with 150-W Lamp and Aeration using $\text{TiO}_2 = 0.2 \text{ g/L}$ pH Trend. ....	42
Figure 39. Flow Through Reactor with 150-W Lamp and Aeration using $\text{TiO}_2 = 0.2 \text{ g/L}$ Ammonia Trend.....	43
Figure 40. Flow Through Reactor with 150-W Lamp and Aeration using $\text{TiO}_2 = 0.2 \text{ g/L}$ Alkalinity Trend. ....	43
Figure 41. Flow Through Reactor with 150-W Lamp with Aeration using $\text{TiO}_2 = 0.2 \text{ g/L}$ COD Trend. ...	44
Figure 42. Absorbance curve of $\text{TiO}_2$ dosing.....	45
Figure 43. Petri dish with $\text{TiO}_2$ and crystal violet.....	46

## LIST OF TABLES

Table 1. USEPA non-hazardous waste landfill effluent limitations (10 CFR 445.21) .....	3
Table 2. Typical leachate water quality data from young and mature landfills (Tchobanoglous, Theisen, and Vigil 1993; Metcalf and Eddy 2003).....	4
Table 3. Summary of extreme values for the composition of leachate developed through review of technical literature.....	5
Table 4. Selected water quality parameters of interest from Dyer Park Landfill leachate (Statom, Thyne and McCray 2004). .....	6
Table 5. Leachate generation rates for 31 Florida landfills (Meeroff and McBarnett 2011).....	8
Table 6. Relative oxidation power of selected oxidizing species (Munter et al. 2001). .....	13
Table 7. Summary of COD removal efficiencies of TiO <sub>2</sub> photocatalytic oxidation from published performance studies. ....	2
Table 8. Measured UV light intensity.....	10
Table 9. Physical composition of Degussa Aeroxide TiO <sub>2</sub> P-25. ....	16
Table 10. Properties of anatase and rutile forms of titanium dioxide. ....	16
Table 11. Falling film reactor experiment using 150-W light source (TiO <sub>2</sub> = 5-g/L).....	25
Table 12. Falling film reactor experiment using 450-W light source (TiO <sub>2</sub> = 5-g/L).....	25
Table 13. Flow through reactor experiment using 150-W light source (TiO <sub>2</sub> = 10-g/L). ....	30
Table 14. Flow through reactor experiment using 450-W light source (TiO <sub>2</sub> = 10-g/L). ....	30
Table 15. Flow through reactor experiment with aeration using 150-W light source (TiO <sub>2</sub> = 10-g/L). ....	35
Table 16. Flow through reactor experiment with aeration using 450-W light source (TiO <sub>2</sub> = 10-g/L). ....	35
Table 17. Flow through reactor experiment with aeration using 150-W light source (TiO <sub>2</sub> = 0.2-g/L). ....	40
Table 18. Alternative AOP test results.....	45
Table 19. Summary of results from using metal additives to stimulate COD removal. ....	46
Table 20: Falling Film UV Light Intensity. ....	48
Table 21: Flow Through UV Light Intensity. ....	48

# ANNUAL REPORT

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**PROJECT TITLE:** Safe Discharge of Landfill Leachate to the Environment

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**KEY WORDS:** Leachate, photocatalytic oxidation, COD, ammonia, alkalinity

**ABSTRACT:** Some closed or partially closed landfills still produce important quantities of leachate, but instead of blending this material with active Class I leachate for disposal, there may be better alternatives. If a relatively inexpensive way to pretreat the leachate and safely dispose of it onsite can be developed, a giant step toward the potential for zero liquid discharge can be achieved. FAU has pioneered the advancement of landfill leachate treatment systems using photochemical iron-mediated aeration and TiO<sub>2</sub> photocatalysis at laboratory scale in previous research funded by the Hinkley Center, which has led to the development of reactor prototypes for pilot scale testing. The objective of the proposed research is to test a prototype photooxidative reactor at pilot scale for the removal of COD/BOD, heavy metals (i.e. arsenic, lead, and iron), ammonia, color, chlorides, and pathogens to determine the feasibility of safely discharging or reusing this leachate as a resource on-site.

**PROJECT WEBSITE:** <http://labees.civil.fau.edu/leahcate.html>

## METRICS:

1. List graduate or postdoctoral researchers **funded** by **THIS** Hinkley Center project.

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Dacey, Justin	CEGE	Meeroff	FAU

3. List research publications resulting from **THIS** Hinkley Center project.

None so far

4. List research presentations resulting from **THIS** Hinkley Center project

TAG meeting presentations only so far

5. How have the research results from **THIS** Hinkley Center project been leveraged to secure additional research funding?

Year two funding from the Hinkley Center for Solid and Hazardous Waste Management was secured. Additional funding is being pursued through the Solid Waste Authority of Palm Beach County and Waste Management.

6. What new collaborations were initiated based on **THIS** Hinkley Center project?

F. Bloetscher, Ph.D., P.E., LEED AP and Craig K. Jones, P.G.

7. How have the results from **THIS** Hinkley Center funded project been used by FDEP or other stakeholders?

To date, the results have not been used by stakeholders yet.

# 1. INTRODUCTION

## 1.1 BACKGROUND

The Hazardous and Solid Wastes Amendments of 1984 required treatment of all surface water runoff from landfills. This act was the first to address the treatment of water from landfills. It was amended in 1991 so that landfills were required to protect groundwater by employing a bottom liner and leachate collection system. The leachate is liquid waste that seeps through the landfill, and any precipitation or other water that comes in contact with waste after collection, must be treated prior to discharge.

The amount of leachate produced from Class 1 landfills in Florida (FDEP 2002; Meeroff and Teegavarapu 2010) can be up to 7000 gallons per day per acre, which must be eventually discharged back into the environment. Landfill leachate is characterized by elevated levels of recalcitrant organic material, color, ammonia, chlorides, and heavy metals such as arsenic, lead, and iron, to name a few. A major limitation to the sustainable management of landfill leachate has been the lack of effective methods to guarantee safe long-term discharge back into the natural environment. This is further complicated by the extremely variable water quality and generation rates, along with the ever-changing regulatory environment, which has caused many conventional technologies to fail in this goal.

In some partially closed landfills, leachate generation rates can be on the order of 400-700 gallons per day per acre (Eyeington 2012). This mature leachate is of very different quality, which is characterized as more stable with a lower BOD<sub>5</sub>/COD ratio, a lower overall organic content, and relatively low ammonia levels (Amokrane et al. 1997; Renou et al. 2008 and several others). One possible way to reduce costs and energy requirements at closed or partially closed facilities would be to treat the mature leachate on-site and reuse or reclaim the water. The first order of business is to determine the water quality parameters of concern when dealing with leachate.

## 1.2 WATER QUALITY REGULATIONS

The quality of the landfill leachate needs to be analyzed so that the most effective treatment process can be designed. As with domestic and industrial wastewater treatment systems, each one is unique, having different constituents that require specialized treatment. The USEPA sets limits to the effluent discharge for non-hazardous sanitary landfills in 40 CFR 445.21 (Table 1). These limits are the very minimum discharge standards, and states and local jurisdictions can apply more stringent limits where applicable.

**Table 1. USEPA non-hazardous waste landfill effluent limitations (10 CFR 445.21)**

<b>Regulated Parameter</b>	<b>Units</b>	<b>Maximum Daily</b>	<b>Maximum Monthly Average</b>
BOD	mg/L as O <sub>2</sub>	140	37
TSS	mg/L	88	27
Ammonia	mg/L as N	10	4.9
α-Terpineol	mg/L	0.033	0.016
Benzoic acid	mg/L	0.12	0.071
p-Cresol	mg/L	0.025	0.014

Regulated Parameter	Units	Maximum Daily	Maximum Monthly Average
Phenol	mg/L	0.026	0.015
Zinc	mg/L	0.20	0.11
pH	Standard units	6-9	6-9

Other limits that are important to note are the USEPA Primary and Secondary Drinking Water Standards, since Florida obtains most of its drinking water from the Floridan or Biscayne aquifers, which are shallow aquifers with direct connections to surface water. Additionally, the Florida Administrative Code F.A.C. 62-550 sets the Florida-specific drinking water standards that must be met. Any landfill that was used prior to the 1976 enactment of the Resource Conservation and Recovery Act (RCRA) may need to apply the F.A.C. 62-777, because before then there was no separation of hazardous and non-hazardous waste, and even after the 1970 Resource Recovery Act, small quantity generators were allowed to dispose of their hazardous wastes in non-hazardous landfills. F.A.C. 62-777 gives specific suggestions for acceptable discharge levels for arsenic, petroleum, benzene, chlorine, and many others.

### 1.3 LEACHATE WATER QUALITY

Several excellent reviews have been conducted with the goal of documenting leachate composition according to the location (i.e. the climate and especially the precipitation rate), the age of the landfill, or the type of wastes. Typical values for selected constituents are summarized in Table 2.

**Table 2. Typical leachate water quality data from young and mature landfills (Tchobanoglous, Theisen, and Vigil 1993; Metcalf and Eddy 2003)**

Constituent	Units	Young	Mature
Ammonia-nitrogen	mg/L as NH <sub>3</sub> -N	10 – 800	20 – 40
BOD <sub>5</sub>	mg/L as O <sub>2</sub>	2000 – 30,000	100 – 200
COD	mg/L as O <sub>2</sub>	3000 – 60,000	100 – 500
Iron (Fe)	mg/L	50 – 1200	20 – 200
pH	pH units	4.5 – 7.5	6.6 – 7.5
Alkalinity	mg/L as CaCO <sub>3</sub>	1000 – 10,000	200 – 1000
TSS	mg/L	200 – 2000	100 – 400

Other important constituents include: i) dissolved organic matter from methane (CH<sub>4</sub>) to volatile fatty acids (VFA) to more refractory humics and fulvics; ii) inorganic constituents, such as calcium (Ca<sup>2+</sup>), magnesium (Mg<sup>2+</sup>), sodium (Na<sup>+</sup>), potassium (K<sup>+</sup>), ammonium (NH<sub>4</sub><sup>+</sup>), iron (Fe<sup>2+</sup>), manganese (Mn<sup>2+</sup>), chloride (Cl<sup>-</sup>), sulfates (SO<sub>4</sub><sup>2-</sup>) and bicarbonates (HCO<sub>3</sub><sup>-</sup>) with heavy metals (arsenic, cadmium, chromium, cobalt, copper, lead, mercury, nickel and zinc), in the microgram per liter range; iii) xenobiotic organic compounds from domestic and industrial sources, comprised of a broad variety of aromatic hydrocarbons, phenols, endocrine disrupting compounds (EDCs), pharmaceuticals, personal care products, pesticides, and chlorinated aliphatics among others.

A review of leachate water quality from 125 sources reported in the literature is summarized in Table 3. The large ranges reported are the result of the high variability among leachates. It is important to note that leachate can have very high concentrations of many different constituents, many of which are known to have deleterious impacts in groundwater and soil. Aside from those listed here, there are numerous other constituents found in leachate ranging from heavy metals (e.g., cadmium, chromium, mercury, arsenic, nickel, selenium, iron, manganese, silver, copper, lead, thallium, zinc and others), other inorganic components (e.g., ammonium, barium, beryllium, bicarbonate, chloride, magnesium, manganese, nitrate, phosphorus, potassium, sodium, sulfate and others) (Qasim and Chiang 1994), and an array of organic constituents including xenobiotic organic compounds (XOCs) such as: BTEX (benzene, toluene, ethylbenzene, and xylene), antibiotics and other pharmaceuticals, pesticides, herbicides and endocrine disrupting compounds (EDCs) (Baun et al. 2003). Specific conditions are not indicated here, as the summary serves to point out the wide variety of leachate water quality that can be found.

**Table 3. Summary of extreme values for the composition of leachate developed through review of technical literature.**

Parameter	Units	Concentration	
		Range	Average
Ammonia	mg/L as NH <sub>3</sub> -N	BDL* – 13,000	1,100
BOD <sub>5</sub>	mg/L as O <sub>2</sub>	BDL* – 80,800	3,100
COD	mg/L as O <sub>2</sub>	0.4 – 152,000	8,750
Conductivity	µS/cm	5.2 – 95,000	15,400
Lead (Pb)	mg/L	BDL* – 5.0	0.41
pH	pH units	2.0 – 11.3	7.73
TDS	mg/L	0.1 – 88,000	11,100
TSS	mg/L	10 – 45,000	1,120
Alkalinity	mg/L as CaCO <sub>3</sub>	3,300 – 11,000	9,640
Color	Platinum-Cobalt Units	3,530 – 40,000	3,630

**BDL\* = below detection limit.**

**Sources:** Adapted from Abu Amr and Aziz (2012), Adlan et al. (2011), Åkesson and Nilsson (1997), Al-Yaqout et al. (2005), Amokrane et al. (1997), Anglada et al. (2011), Aziz et al. (2011), Bashir et al. (2010), Bekbölet et al. (1996), Bernard et al. (1997), Bila et al. (2005), Bouhezila et al. (2011), Calli et al. (2005), de Moraes and Zamora (2005), Deng and Ezyske (2011), Geenens et al. (2000), Gonze et al. (2003), Hickman (2003), Iaconi et al. (2010), Imai et al. (1998), Ince (1998), Jia et al. (2011), Kim et al. (1997), Kim et al. (2007), Kjeldsen et al. (2002), Kurniawan and Lo (2009), Li et al. (2009), Lin et al. (2000), Mahmud et al. (2011), Mohammad et al. (2004), Mohajeri et al. (2010), Moraes and Bertazzoli (2005), O’Leary and Walsh (1995), Oweis and Kehra (1998), Poblete et al. (2012), Tammemagi (1999), Tatsi et al. (2003), Tchobanoglous and Kreith (2002), Reinhart and Grosh (1998), Reinhart and Townsend (1998), Renou et al. (2008), Salem et al. (2008), Silva et al. (2004), Solid Waste Authority of Palm Beach County (2006), Statom et al. (2004), Steensen (1997), Tamrat et al. (2012), Vilar et al. (2011), Ward et al. (2002), Westlake and Phil (1995), Wichitsathian et al. (2004), Wu et al. (2004), Youcai et al. (2002), Zhao et al. (2010).

It is difficult to define a typical landfill leachate water quality because each landfill produces varying compositions of leachate at different times depending on waste composition, climate, seasonal variations, rainfall, and solid waste management practices. In this regard, each landfill

is unique. In this particular study, the Dyer Park Landfill located in Palm Beach County, Florida is the focus because it is a partially closed landfill that generates particularly weak leachate. The Dyer Park landfill operated from 1968 to 1992. However, the lined section of the landfill accepted waste from 1984 to 1992. A 2004 study (Statom, Thyne and McCray 2004) investigated the leachate water chemistry of the Dyer Park landfill and monitored the levels of several contaminants, as summarized in Table 4.

**Table 4. Selected water quality parameters of interest from Dyer Park Landfill leachate (Statom, Thyne and McCray 2004).**

Parameter	Units	No. of Samples	Range	Mean	Standard Deviation
pH	Standard units	50	6.56 – 8.01	7.07	0.27
Conductivity	mmhos/cm	49	3.6 – 15	7.64	2.85
Temperature	°C	50	2.16 – 32.8	27.7	2.14
COD	mg/L	50	222 – 2000	835	383
BOD	mg/L	48	<1 – 184	47	40.2
Ammonia	mg/L	50	5.6 – 1350	473	254
Chloride	mg/L	49	63 – 1580	837	330
Sulfate	mg/L	49	<1 – 118	20	26
Alkalinity	mg/L	31	1160 – 3900	2450	597
Bicarbonate	mg/L	4	1900 – 3900	2660	928
Calcium	mg/L	23	132 – 220	176	22
Magnesium	mg/L	20	41 – 63	54	5.7
Arsenic	µg/L	49	<5 – 25	nr	nr
Iron	mg/L	50	1.6 – 9.7	4.8	2.4
Boron	mg/L	6	2.6 – 4.0	3.2	0.5
Chromium	µg/L	49	<5 – 60	20	11
Lead	µg/L	49	<4 – 110	nr	nr
Silver	µg/L	49	<1 – 25	nr	nr
Zinc	µg/L	49	<6 – 488	nr	nr

#### 1.4 POLLUTANTS OF CONCERN

This research will focus on the following parameters: chemical oxygen demand (COD), alkalinity, ammonia, color, and pathogens. These constituents are some those typically detected in the Dyer Park Landfill leachate, currently there is no treatment process being used, the leachate is disposed of via deep well injection.

COD was chosen as a measure of the oxygen demand of both organic and inorganic matter in the leachate (USEPA 2012). High levels of COD are caused by the low levels of dissolved oxygen found inside the landfill and the leachate. High concentrations of COD (>400 mg/L) can be devastating to aquatic life by removing all oxygen in the receiving water and causing anoxic conditions leading to asphyxia. The mean value found in 2004 was 835 mg/L. There is no specific treatment target for COD, but the COD that causes dissolved oxygen levels to get below 3 mg/L will cause adverse aquatic water quality impacts since the ideal range of dissolved

oxygen is 5 mg/L or higher for aquatic life. The federal government currently has set limitations for BOD (see Table 1), but not for COD. Local sewer use limitations (for example Broward County Code Chapter 34 Article VI, Ordinance No. 2001-43 Sewer Use Ordinance) typically charge a fine for high strength wastewater if COD concentration exceeds 800 mg/L.

Alkalinity is the amount of protons a solution can accept without a measurable change in pH (Tchobanoglous and Kreith 2002). Alkalinity allows the pH range to resist change with the input of acidic or alkaline agents. The pH range that is typical for water containing significant amounts of alkalinity is between 6.0 and 8.5. The typical levels of alkalinity for Dyer Park Landfill leachate average 2,450 mg/L with the majority being comprised of bicarbonate. While effects on humans have been found with high levels of alkalinity, the clogging of pipes can effect distribution networks and discharge structures.

Ammonia ( $\text{NH}_3$ ) is an inorganic form of nitrogen that is created in the natural anaerobic degradation process of many organic compounds (USEPA 2012). When ammonia in an aqueous solution is exposed to air, it rapidly becomes a colorless gas with a strong noticeable odor (Commonwealth of Australia 2010). The amount of ammonia in an aqueous solution has a direct correlation to temperature and pH (Environmental Protection Agency 1985). Concentrations of ammonia at levels of 0.03 mg/L have been found to be toxic to aquatic life, and the  $\text{LC}_{50}$  (Lethal Concentration which is fatal to 50% of the subjects) for freshwater fish occurs at 0.068 – 2.00 mg/L as  $\text{NH}_3\text{-N}$ , during a set exposure time of 96 hours (Eddy 2005). Taste and odor issues have been reported at levels of 35 mg/L and 1.5 mg/L, respectively (WHO 2004). Ammonia concentrations naturally found in groundwater and surface water are usually less than 0.2 mg/L, although anaerobic groundwater may have levels near 3 mg/L. In the state of Florida, ammonia is identified as a “minimum criteria systemic toxicant” and has a groundwater cleanup target level (CTL) of 2.8 mg/L. The CTL is not a regulation or standard, but rather a suggestion for water quality. The concentrations found in leachate, which were shown in Table 2 (20-40 mg/L for mature leachate) and Table 3 (up to 13,000 mg/L as  $\text{NH}_3\text{-N}$ ) far exceed these levels. Broward County sewer use limitations stipulate high strength wastewater surcharges if the  $\text{NH}_3\text{-N}$  is above 25 mg/L as  $\text{NH}_3\text{-N}$ , and concentrations exceeding 70 mg/L as  $\text{NH}_3\text{-N}$  are not permissible. The target level of ammonia is 4.9 mg/L from EPA 40 CFR 445.21. For comparison purposes, the Dyer Park Landfill leachate historically averages 473 mg/L of ammonia.

While color has no effect on the quality of water, it is one of the most noticeable of all its traits. The standard water color set by the USEPA is 15 Platinum-cobalt units (USEPA 2012). With respect to pathogens, which can affect human health, the lowest amount of pathogens should be discharged from the landfill. Further study is needed regarding this issue in leachate, but the basis for leachate discharge should conform to the drinking water standard of the “*no more than 5% of the monthly tests*” should test positive for total coliform. This will protect all of the recreational users of the waterway, and not but undo stress on the environment.

## 1.5 LEACHATE QUANTITIES

The volume of leachate generated depends on the amount of rain that percolates through the landfill and the exposed surface area. Other factors that influence the volume of leachate include: surface runoff, groundwater intrusion, liquid waste in the landfill, irrigation, evapotranspiration,

landfill depth, and refuse composition (Westlake and Phil 1995), but the quantity of leachate is directly tied to the amount of precipitation and irrigation that the landfill receives.

A survey was performed by Florida Atlantic University (Meeroff and Teegavarapu 2010) polling 52 landfills in the state of Florida about their leachate generation rates. Facilities were divided into four different size classes based on their waste capacity: small (500,000 MT/yr), medium (5,000,000 MT/yr), large (15,000,000 MT/yr), and super (> 15,000,000 MT/yr) (USEPA 1999). The results of the survey from the 31 facilities that responded showed leachate volumes ranging from less than 100 to nearly 3,000 gpd/acre (refer to Table 5).

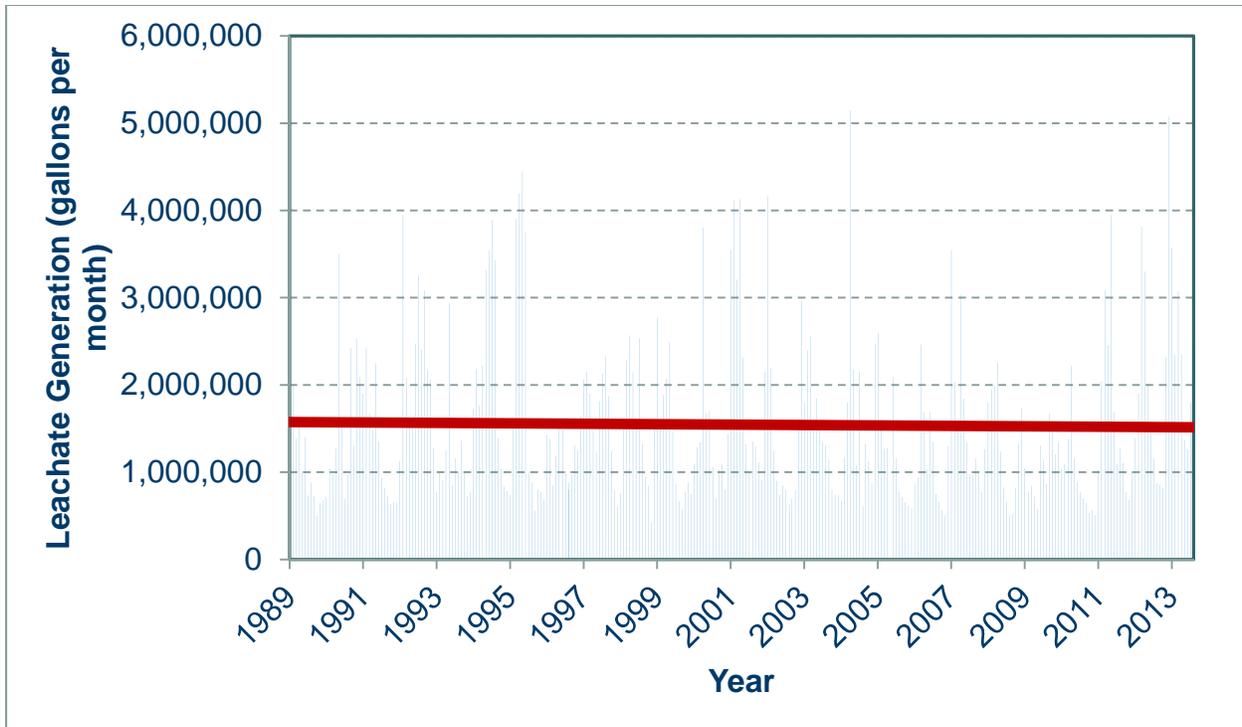
**Table 5. Leachate generation rates for 31 Florida landfills (Meeroff and McBarnett 2011).**

Class	Range (gpd/acre)	Number of landfills
Small	<100	14
Medium	100-300	9
Large	300-850	6
Super	>850	2

The Hydrologic Evaluation of Landfill Performance (HELP) model is a computer program developed by the Waterways Experiment Station (WES), which is the headquarters for the U.S. Army Engineer Research and Development Center (ERDC) (USEPA 2012). The HELP model is used to estimate the generation of leachate from landfills for comparison efforts in the planning and design of the landfill and leachate collection system. The HELP model gives a theoretical value in the South Florida area of 2,000 – 3,000 gpd/acre, which is the design value used for most landfills in the Southeast Florida region. However, most landfills do not have properly calibrated flow meters to record the actual leachate volumes and leachate is also generated in partially lined cells or older systems (Meeroff and McBarnette 2011), so accurate generation values are not readily available.

The Dyer Park Landfill is a partially capped landfill in that it has a cover over the top flat surface but the side slopes are uncapped. So any irrigation or precipitation that falls on the slope could penetrate into the cell and eventually become leachate. According to historical data, the quantity of leachate generated from the Dyer Park landfill comprises approximately 10-25% of the overall leachate flow for the SWA facility. This is on the order of 600,000 – 5,000,000 gallons per month.

The eastern section of Palm Beach County has an average rainfall of 62 inches of rain per year (Statom, Thyne and McCray 2004). Thus, over the 80-acre landfill area, the amount of leachate generated should average 200,600 gallons per month, but data provided by the Solid Waste Authority of Palm Beach County shows a historical average flow of 1.5 million gallons a month with a peak flow of 5 million gallons/month, seen in Figure 1.



**Figure 1: Dyer Park Historical Leachate Quantities (SWA personal communication 2014).**

To put these amounts in perspective, Palm Beach County residents used an average of 50-million gallons of water a day in 2011 (Palm Beach County 2011). That means that every 2.7 years, the daily water usage of Palm Beach County is disposed of in a deep well, eliminating a reusable resource in a non-sustainable manner.

## **1.6 METHODS OF LEACHATE MANAGEMENT**

As a consequence of collecting these concentrated volumes of leachate, containing synthetic organic compounds, heavy metals, and other constituents of concern as discussed earlier, the liquid waste must be eventually discharged back into the environment. A major limitation to the sustainable management of landfill leachate has been the lack of effective methods to guarantee safe long-term discharge back into the natural environment. This is further complicated by the extremely variable water quality and waste generation rates, along with the ever-changing regulatory environment, which has caused many conventional technologies to fail in this goal.

Currently viable leachate management options include: on-site treatment, municipal sewer discharge, deep well injection, hauling offsite, or a combination approach. In the case that deep wells cannot be permitted or hauling is not cost effective, municipal sewer discharge is favored. However, wastewater treatment plants are facing the possibility of having to meet discharge limits (for nutrients and emerging contaminants of concern) that exceed the boundaries of current technologies. Facilities that accept leachate may struggle to meet the proposed new limits (i.e. USEPA numerical nutrient criteria), and may stop accepting the material or impose excessively high surcharges. So it is conceivable that in the future, municipal sewer discharge will become a

limited option. Given this scenario, landfill managers may soon be forced into on-site treatment to handle their leachate.

In previous work funded by the HCSHWM, 23 different engineering alternatives for long-term leachate management were evaluated (Meeroff and Teegavarapu 2010). For on-site treatment to work, some form of aerobic treatment would be expected to reduce leachate strength prior to discharge. However, biological systems are not well-suited for the removal of bio-toxics from water and are inefficient in dealing with wastes of varying quality, such as leachate. Thus post-treatment, using constructed wetlands, combined physicochemical treatment, or evaporation systems, would then be required. Unfortunately, technologies such as activated carbon and certain advanced treatment processes, such as ozone, do not adequately address inorganics, and membrane systems or air stripping merely transfer organics to another phase or create a side stream, like concentrate brine, that cannot be discharged readily. Furthermore, multiple barrier systems are complicated to operate, costly, and generally inefficient. For on-site treatment options, the most effective strategies involve technologies that can destroy different classes of harmful contaminants all at once, without producing adverse byproducts and residuals.

There are basically only a few major ways of treating or disposing of leachate. Treatment can be off-site transfer with or without pre-treatment or reused on-site with or without pre-treatment in the form of biodegradation, physicochemical treatment, or a combination approach. The current method used at Dyer Park is deep well injection, which is disposal without treatment.

One viable option for landfill managers is hauling off-site. Landfills will collect their leachate and send truckloads of the liquid waste to an ultimate disposal site; typically an off-site publicly-owned treatment works (POTW), where it is combined with domestic wastewater and processed along with the municipal sewage. This method does not address the ultimate disposal of leachate; it simply moves the leachate to another location off-site. The option presents a high transportation risk and can be a potentially expensive solution, depending upon the distance to the receiving site and the treatment performance of the facility accepting the material. If the travel distance is relatively short (<100 miles), the costs can be very competitive, and this can be a viable option. But if the site is located at distances greater than 100 miles, the costs can be potentially limiting. For example, Polk County, FL reported a three-year contract they signed in July 2009 for the disposal of their landfill leachate at \$130 per thousand gallons (Blandford 2011). Another issue is that not all wastewater treatment plants will accept leachate due to the extremely high concentrations found. In addition, leachate generally has low biodegradability and may contain heavy metals. So large volumes of leachate can upset the normal biological treatment processes at the POTW (Boyle and Ham 1974; Booth et al. 1996), which may lead to expensive surcharge rates or even rejection. The costs associated with hauling can also vary depending on: the cost of fuel, the distance the leachate needs to travel, and if there may be a need to pre-treat the liquid waste prior to wastewater treatment plant acceptance. Besides the transportation risk and fuel cost volatility, the most problematic issue is if the contractor at the facility accepting the leachate suddenly decided that the material is not profitable to handle, treat, and dispose of safely, and terminated the agreement to accept the leachate. This will become particularly problematic if regulations governing wastewater disposal were to become more stringent with respect to ammonia-nitrogen, toxic trace metals, and/or inhibitory organic compounds with low biodegradability. The facility accepting the waste may find that leachate volumes are too high (Çeçen and Çakıroğlu 2001) compromising the treatment plant's ability to

meet permitted discharge water quality levels. In this case, the treatment facility would likely consider no longer accepting the material.

Leachate recirculation is another option for managing leachate. This process consists of reintroducing the landfill leachate back into the landfill. The recirculating leachate accelerates the breakdown of organic materials within the landfill (Xing et al. 2012). This leads to increased methane production which must be managed to prevent elevated photochemical ozone formation (Xing et al. 2012). The build-up of head pressure from the increased amount of leachate in the bottom of the landfill creates higher potential for the leachate to escape the landfill into the environment and towards the ground water and soil. Tropical climates make leachate circulation difficult due to high temperatures and elevated levels of evaporation, which lowers the moisture content of the solid waste thereby diminishing the biological activity. One study determined the effect of applying additional water in order to maintain certain levels of moisture, on the methane production and stabilization of the landfill. It was found that supplementing the leachate with water kept elevated levels of methane production and lower time periods of landfill stabilization (Sanphoti et al. 2006). Recirculation can improve moisture content and distribute nutrients and enzymes throughout the landfill. The COD of the leachate was lowered with this process but it also lowered the methane production of the landfill (Chugh, et al. 1998). The implementation of a leachate recirculation system requires high capital and recurring maintenance costs. Also, odor problems from leachate recirculation are common (McBarnette 2011).

If the appropriate aquifer conditions exist and permitting is available, another attractive disposal option for leachate is deep well injection. Essentially, this is the same as transferring the leachate off-site without treatment. In this option, the leachate is pumped deep into the ground below the aquifer and between confining layers to assure separation from the groundwater supply. The biggest concern with deep well injection is the risk of contamination of the drinking water supply (Groundwater Protection Council 2005). The exact geology thousands of feet underground can be difficult to establish with complete certainty. Even a minor fracture can cause a substantial problem as groundwater remediation is an incredibly difficult task at these depths. On July 7, 2000, the United States Environmental Protection Agency (EPA) proposed revisions to the underground injection control (UIC) regulations that would restrict wastewater injection by existing Class I municipal disposal wells that have caused or may cause movement of contaminants into underground sources of drinking water (USDWs) in specific areas of Florida (65 FR 42234) unless the owner met certain additional requirements: 1) advanced wastewater treatment (AWT) and high level disinfection (so that primary health-based drinking water standards would not be violated) with a non-endangerment demonstration or 2) in-depth hydrogeological demonstration and advanced treatment, as necessary. The second option refers to 40CFR146.15(d), which states that to qualify for authorization the owner shall develop and implement a pretreatment program that is no less stringent than the requirements of Chapter 62-625, Florida Administrative Code, or have no significant industrial users as defined in that chapter. Furthermore, the owner must treat the injectate using secondary treatment in a manner that is no less stringent than the requirements of Florida Rule 62-600.420(1)(d), and using high-level disinfection in a manner that is no less stringent than the requirements of Florida Rule 62-600.440(5)(a)-(f). The specified treatment requirements then are designed to achieve an effluent after disinfection containing not more than 20 mg/L CBOD<sub>5</sub> and 20 mg/L TSS, or 90% removal of each of these pollutants from the wastewater influent, whichever is more stringent. The

bottom line is that the proposed rule requires installation of advanced wastewater treatment with high level disinfection by 2015 for Class 1 injection wells. This means that injected water would need to meet at a minimum, secondary treatment and high-level disinfection as defined in the Florida regulations, with filtration required for total suspended solids (TSS) control prior to disinfection (such that the treated wastewater contains no more than 5.0 mg/L of TSS before the application of the disinfectant). Existing (installed before December 22, 2005) Class I municipal disposal wells that have caused or may cause fluid movement into USDWs in specific counties and under certain geologic conditions in Florida are limited to: Brevard, Broward, Charlotte, Collier, Flagler, Glades, Hendry, Highlands, Hillsborough, Indian River, Lee, Manatee, Martin, Miami-Dade, Monroe, Okeechobee, Orange, Osceola, Palm Beach, Pinellas, St. Johns, St. Lucie, Sarasota, and Volusia counties. These counties were targeted in the proposal because they have the unique geologic conditions that are predominated by carbonate rocks. Class 1 industrial disposal wells are designed with additional levels of protection, and it is unclear whether industrial disposal wells have suffered the same issues as municipal disposal wells. The basic regulations for industrial wells are not substantially different than municipal wells, but an additional layer of protection is designed and built into industrial wells. The only differences include seismic studies for structural integrity, an additional area of review of 2 miles compared to 0.25 miles, modeling studies to show the waste will remain in the injection zone for 10,000 years, automatic alarms and shutdown devices, year internal mechanical integrity testing, and additional monitoring and reporting requirements. They are built with a tubing and packer system that places the injection tube inside a liquid filled annulus. This allows for better monitoring of performance.

The unique problems associated with deep well injection are not limited to the solid waste manager owning an onsite well. Many of the large publically-owned treatment works (POTWs) in those same counties dispose of treated wastewater effluent via deep wells. Therefore, if the leachate compromises the ability to meet the discharge limits set forth in the injection permit, the POTW may not wish to accept the leachate. This would cause problems for another important leachate management option of municipal sewer discharge.

As the options begin to get limited, on-site pretreatment becomes more and more necessary. Typically, some form of aerobic biological treatment would be required to reduce leachate BOD strength prior to discharge. Biodegradation is performed by microorganisms, which degrade organic compounds under aerobic conditions and convert soluble BOD into particulate BOD, which can be readily removed via sedimentation. Anaerobic digestion can also work for dealing with strong leachate but will have long treatment times (Berrueta and Castrillon 1993). Nevertheless, biological systems are not well-suited for removal of biotoxics from water and are inefficient in dealing with wastes of varying quality, such as leachate. Thus post-treatment, using constructed wetlands, combined physical/chemical/biological treatment, or evaporative systems, is generally required (Booth et al. 1996). Unfortunately, activated carbon and certain advanced treatment processes, such as ozone or ultraviolet light, do not adequately address inorganics, and membrane systems or air stripping merely transfer organics to another phase. Furthermore, multiple barrier systems are complicated to operate, costly, and generally inefficient. Unfortunately, most current processes cannot adequately address inorganics and organics simultaneously. From our previous work funded by the HCSHWM (Meeroff, Gasnier, and Tsai 2006; Meeroff, Gasnier, and Tsai 2008; Meeroff and Teegavarapu 2010), our research team

evaluated 23 different engineering alternatives for long-term leachate management. The results indicated that the most effective and sustainable strategies for the future would involve technologies that can destroy different classes of harmful contaminants all at once, without producing adverse byproducts and residuals.

Physical and chemical treatment options include floatation, coagulation/flocculation, adsorption, air stripping and chemical oxidation. Coagulation/flocculation is commonly used as a pre-treatment to remove non-biodegradable organic matter using aluminum sulfate (Amokrane, Cornel and Veron 1997). The disadvantage is the large volumes of chemical sludge produced. Flocculation can also foul filters in the treatment process, or require multiple filters to prevent this clogging. Chemical oxidation is widely studied and is of growing interest with a focus on advanced oxidation processes (AOP), which employ strong oxidants and ultraviolet light. The process works best on old, well-stabilized leachate (Renou, et al. 2007).

## 1.7 ADVANCED OXIDATION PROCESSES

Advanced oxidation processes (AOPs) are methods that promote the creation of and utilize highly reactive oxidants such as hydroxyl radicals, ozone, and chlorine (Table 6). These oxidation reactions produce radicals which are chemical species that possess an unpaired electron, causing them to be very unstable. The unstable radicals attempt to stabilize themselves by reacting with surrounding species. The radicals will continue to react with substances until stability is reached. Within milliseconds (Peyton and Glaze 1988, cited by Fang et al. 2005), hydroxyl radicals are capable of achieving complete mineralization (i.e. degradation of complex organics to CO<sub>2</sub>, H<sub>2</sub>O, and mineral ions) of virtually all organic compounds (Feitz et al. 1999; Cho et al. 2002) rather than concentrate or transfer contaminants into a different phase. In this manner, pollutants that are only partially oxidized are decomposed into components that are more readily biodegradable and less toxic to common microorganisms found in a wastewater treatment plant for instance (Schulte et al. 1995; de Morais and Zamora 2005).

**Table 6. Relative oxidation power of selected oxidizing species (Munter et al. 2001).**

Oxidation Species	Symbol	Relative Oxidation Power
Positively charged hole on titanium dioxide	( $h^+$ )	2.35
Hydroxyl radical	(OH·)	2.05
Ozone	(O <sub>3</sub> )	1.52
Hydrogen peroxide	(H <sub>2</sub> O <sub>2</sub> )	1.31
Permanganate	(MnO <sub>4</sub> <sup>-</sup> )	1.24
Hypochlorous acid	(HOCl)	1.10
Chlorine	(Cl <sub>2</sub> )	1.00

The focus of this research is the use of photocatalytic oxidation with titanium dioxide (TiO<sub>2</sub>), in combination with ultraviolet light energy (UV) in the presence of oxygen. TiO<sub>2</sub> is a white semiconducting powder consisting of nanoparticles with an average size of 200-nm. The crystalline structure of TiO<sub>2</sub> is available in three forms: anatase, brookite, and rutile (Ohtani, et al. 2010). TiO<sub>2</sub> was rated as a Group 2B (possibly carcinogenic to humans) substance by the International Agency for Research on Cancer (IARC) in 2006 as an inhalation hazard because of its nature as a fine dust. The health limitations on TiO<sub>2</sub> are only expressed for inhalation, and

there are none listed for ingestion. Furthermore, no adverse effects have been measured to fish, daphnia, or bacteria in ecotoxicological studies. The material is extensively used in products such as paints and varnishes, floor coverings, roofing granules, sunscreens and cosmetics, printer inks, ceramics, plastics, paper coatings, pigments used in numerous foods, toothpastes, medicines, dielectric mirrors and tattoo pigments (Department of Health and Human Services 2011).

TiO<sub>2</sub> particles generate strong oxidizing power when illuminated with UV light at wavelengths less than 400 nm. Irradiation of TiO<sub>2</sub> with photons of ultraviolet light energy produce areas of positive charge in the valence band of the semiconductor (“holes”) and free electrons in the conduction band. When the “holes” and free electrons interact with water trapped in the pores of the catalyst, a mixture of indiscriminate oxidants are generated including hydroxyl radical (HO<sup>•</sup>) and superoxide radical (O<sub>2</sub><sup>•-</sup>). For photocatalysis to occur, electron “holes” must migrate to the surface of the TiO<sub>2</sub> crystal. The “holes” primarily react with hydroxide (OH<sup>-</sup>) from water acting as electron donors to produce hydroxyl radicals (Rincon and Pulgarin 2005). The electrons primarily react with O<sub>2</sub>(aq) (dissolved oxygen) in water acting as electron acceptors to yield the superoxide radical. Some of the electron-hole pairs, which do not participate in the redox reaction with water or oxygen, disappear as heat losses via the recombination of holes and electrons. Utilizing the combined oxidation power of holes and hydroxyl radicals generated in the valence band (VB), and electrons and superoxide radicals generated in the conduction band (CB), illuminated TiO<sub>2</sub> photocatalysts can decompose organic compounds by participating in a series of mineralization reactions (Rincon and Pulgarin 2005). Chong et al. (2010) explained that the oxidative and reductive reactions from titanium dioxide are due to its unique characteristic of possessing a lone electron in its outer orbital. The reaction process begins when UV light energy photoexcites that lone outer shell electron, which creates an empty outer valence band.

Basically, the photocatalytic process is an array of multi-step reactions. The ability of TiO<sub>2</sub> photocatalyst to mineralize a wide range of pollutants is an attractive quality, but modeling the kinetics of such a complicated process is a difficult task. Sometimes, complex environmental processes allow only for empirical solutions because not all reactions or mechanisms are known. There may be lumped parameters, surrogates, indicators or just overly complex reaction pathways (Hemond and Fechner-Levy 2000).

Recently, we have developed a falling pilot film reactor for preliminary pilot testing at FAU and have been able to achieve removal of 34% COD, 57% color, 84% alkalinity, and 82% ammonia within 24 h of treatment at a TiO<sub>2</sub> dose of 4–10 g/L. Although the process did not reduce the COD concentrations to below 800 mg/L, we were able to demonstrate destruction of 1400 – 2400 mg/L of COD in just 24 h. These longterm experiments concluded that first order reaction kinetics best fit the observed destruction of most water quality parameters. So it may be possible to meet the requirements of F.A.C. 62-302 for metals and 62-777 for surface water target levels using this technology and to develop parameters for scale-up. Thus it may now be possible to eliminate impurities in water all at once using a single process, and if these processes work as well in the field (at pilot scale) as they do in the laboratory, then we may have the answer for landfill managers when they run out of options for safely managing their leachate. The questions that remain focus on the suitable intensity of the ultraviolet light radiation to apply and a simple

method to determine the appropriate amount of photocatalyst to use for treatment due to the concentration dependence discovered in previous work. Refinements of the process still need to be worked out with respect to recovering the photocatalyst after a batch of treatment and determining the recovery number, which is related to the number of times the catalyst can be reused before it is spent. These improvements will allow operation at a much lower cost. A summary of research performed using  $\text{TiO}_2$  photocatalysis, focusing on COD removal is listed in

**Table 7.**

**Table 7. Summary of COD removal efficiencies of TiO<sub>2</sub> photocatalytic oxidation from published performance studies.**

Water Type	TiO <sub>2</sub> Dose (g/L)	UV (W)	COD <sub>0</sub> (mg/L)	pH	Removal (%)	Time (min)	Reference
Grey water	2.0-5.0	nr (TQ 150z1)	3940	10.3	44	150	Sanchez et al. 2010
Simulated wastewater	1% Pt-TiO <sub>2</sub> immobilized on silica gel	88 W (1.8 mW/cm <sup>2</sup> )	62	6.5	86	30	Suri et al. 1999
Simulated wastewater	0.3-1.0	8 W	10	n/a	82	120	Huang et al. 2008
Lagoon wastewater	2.0	Solar radiation	660	8.0	42	120	Araña et al. 2002
Industrial wastewater	0.6	6 × 18 W	3.2	6.0	62	60	Chen et al. 1997
Olive mill wastewater	1.0	Solar radiation (assumed 30 W/m <sup>2</sup> )	6,600	2.8	26	1920	Gernjak et al. 2004
Industrial wastewater	4 plates immobilized	4 × 4 W	120 (TOC)	9.0	34	30	Nakamura et al. 2008
Industrial wastewater	1.0	415 W	135	8.0	22 (diluted 1:100 + filtered)	1440	El Hajjouji et al. 2008
Industrial wastewater	0.5	400 W	404	3.0	40	240	Pekakis et al. 2006
Landfill leachate	5.0 (batch) immobilized	16 × 40 W 5.0 – 10.0 mW/cm <sup>2</sup>	985	5.0	70	480	Bekbölet et al. 1996
Landfill leachate	3.0	8 W (21 W/cm <sup>2</sup> )	1,673	8.7	30	720	Cho et al. 2004
Landfill leachate	1.0-2.0	150 W (0.5 mW/cm <sup>2</sup> )	1,200	7.5	35-57	60	Poblete et al. 2012
Landfill leachate	TiO <sub>2</sub> coated sheet	120 W	26,000 – 30,000	5-7.6	76-92	150	Chemlal et al. 2013
Industrial wastewater	3.0	7.6 W/m <sup>2</sup>	20,000	6.8	36.3	1440	Baransi et al. 2012
Waste Activated Sludge	3.0	1.5 mW/cm <sup>2</sup>	16,249	6.83	45	480	Liu et al. 2012

<b>Water Type</b>	<b>TiO<sub>2</sub> Dose (g/L)</b>	<b>UV (W)</b>	<b>COD<sub>o</sub> (mg/L)</b>	<b>pH</b>	<b>Removal (%)</b>	<b>Time (min)</b>	<b>Reference</b>
Landfill leachate	2.0	NA	2,440	8.24	60	4320	Jia et al. 2013
Simulated wastewater	3.2 g of TiO <sub>2</sub> coated on immobilized sheet	38 W/m <sup>2</sup>	157,000	7.0	51.6	255	Yahiat et al. 2011
Paper mill wastewater	0.75	35-45 W/m <sup>2</sup>	2,075	6.5	70.5	180	Ghaly et al. 2011
Industrial wastewater	0.2	Solar radiation	500	6	32	240	Vineetha et al. 2012
Landfill leachate	200 mg/L	Solar radiation	3270 to 4575	5	63%	1019 kJ/L	Rocha et al. 2011
Landfill leachate	2 to 10 g/L	15 W	2440	8.2	60%	4320	Jia et al. 2012
Swine and bovine manure	Immobilized with 1 g/L suspension	6-100 W UVA	Ammonia 100 ppm	7.1-8.7	55%	360	Altomare et al. 2012

A primary advantage of photocatalytic technologies over other advanced oxidation processes is that photons in the near-UV range can be used instead of the dangerous radiation of the UV-C region required for UV disinfection or oxidation. Consequently, energized processes can potentially make possible the use of free sunlight instead of expensive mercury lamps (Bolduc and Anderson 1997). To date, experiments conducted at FAU have used less than 2 mW/cm<sup>2</sup> of ultraviolet energy, which is 50 times less than natural sunlight (100 mW/cm<sup>2</sup>). The technology is easy to operate because the process just requires sufficient contact time and does not rely on complex precipitation reactions, chemical addition, or biochemical processes. Another major advantage is the simultaneous removal of organics, metals, and pathogens without merely transferring the pollutant to another medium (i.e. air or sludge). Therefore, this technology may provide an efficient, environmentally-friendly, and sustainable approach to long-term leachate management as well as aquatic water quality protection. Potential applications extend beyond solid waste management and include indirect potable reuse, water recycling, aquifer recharge, advanced wastewater treatment, and even household or portable systems. These methods will allow landfill operators with little training to reliably manage leachate without spending too much time on the task.

## **1.8 OBJECTIVES**

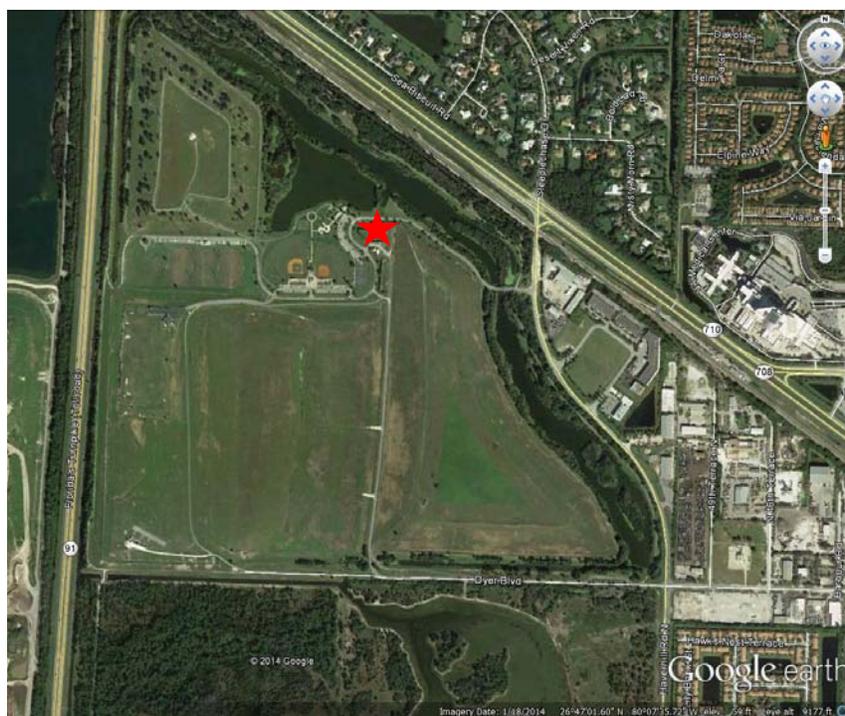
The main purpose of this research is to test UV/TiO<sub>2</sub> photocatalytic degradation of selected pollutants (COD, ammonia, alkalinity, and color) in mature landfill leachate using a pilot scale falling film reactor. It is hypothesized that the degradation rate of the contaminants will follow overall first-order reaction kinetics. The primary objective of this study is to determine an optimum dosage of catalyst that meets the water quality goals of one or more of the following: 1) surface water discharge, 2) industrial reuse as cooling water or irrigation, or 3) on-site use as dilution water to reduce leachate clogging issues in pipes.

## 2. METHODOLOGY

### 2.1 LEACHATE COLLECTION

Leachate for experimental testing was collected from the Dyer Park Landfill located in Palm Beach County, FL. The Dyer Park Landfill is operated by the Solid Waste Authority (SWA) of Palm Beach. The landfill is currently partially closed, no longer accepting waste, and is used as recreational facility. The footprint of the landfill is approximately 80-acres. The leachate collection system is operated by SWA, and the facility collected un-milled municipal solid waste from 1984 to 1992 for disposal.

Leachate was sampled with permission from Nathan Mayer P.E. Samples were collected from the pump station located on the northwest corner of the landfill, designated by the star on Figure 2.



**Figure 2. Dyer Park Landfill sample collection point.**

The first leachate samples were collected on May 30, 2014. The leachate was collected from the Dyer park leachate pumping station (Figure 3). The samples were taken from a ¼-inch sampling port with valve (Figure 4). The sampling port was purged for one minute before collection started (Figure 5). The leachate sample was collected in a five-gallon HDPE container (Figure 6). Measurements of dissolved oxygen, pH, temperature, and conductivity were taken directly after collection using a YSI 550MPS, as described later. The sample was placed in a cooler to preserve the sample and limit its exposure to light. The sample was stored in a refrigerator at 4°C until treated in the laboratory.



**Figure 3. Dyer Park sampling point pump station dry well.**



**Figure 4. Dyer Park sampling port at the pump station.**



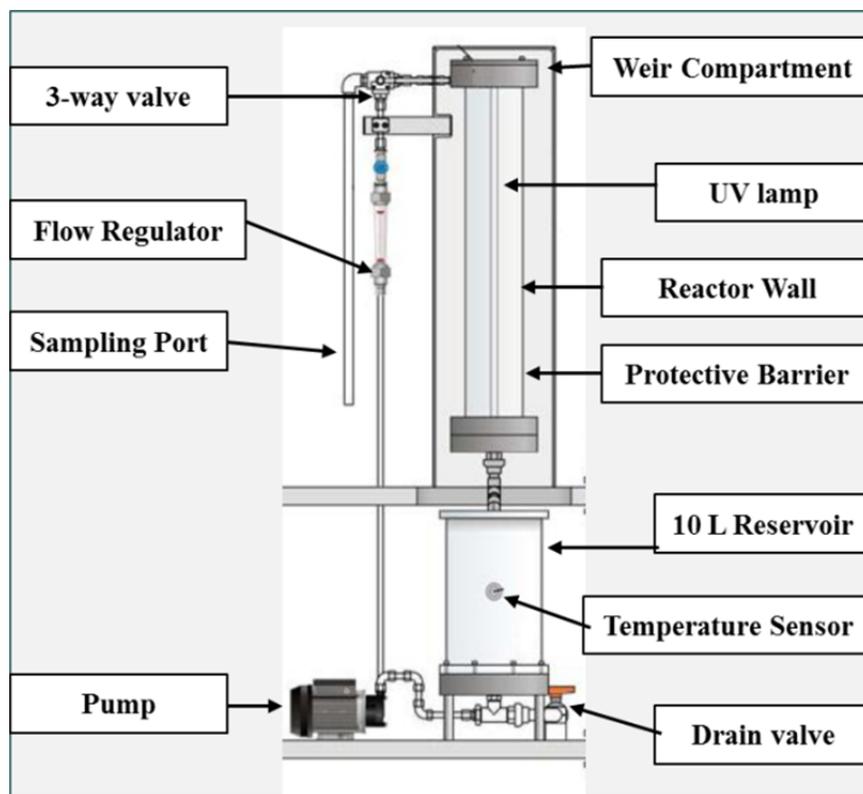
**Figure 5. Purging the sample port prior to collection.**



**Figure 6. Filling 5-gallon sample containers.**

## **2.2 PILOT SCALE REACTOR**

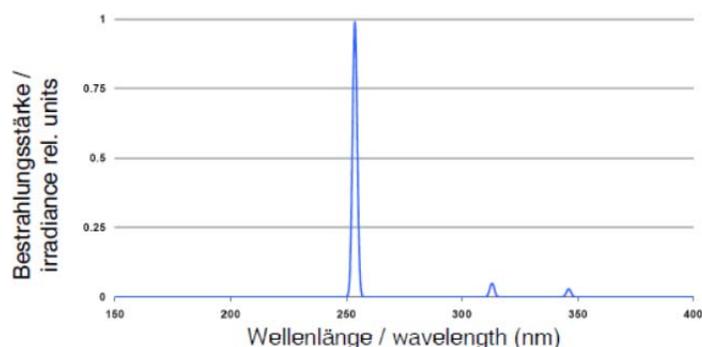
The experiments for this pilot scale research were conducted using a: CE 584 Advanced Oxidation, which is part of the 2E – Energy and Environment product range. 2E is a sector owned by G.U.N.T. Gerätebau GmbH; a company based in Barsbuettel, Germany. The advanced oxidation unit is shown in Figure 7. The reactor was operated in two configurations a falling film reactor and a flow through reactor. The reactor measures 1510 mm × 790 mm × 1990 mm and weighs approximately 330 lbs. The main components of the advanced oxidation unit are labeled in Figure 7.



**Figure 7. Main components for the pilot scale reactor.**

The unit is equipped with a 10-L reservoir, temperature sensor (0-50°C), 260 Liter per hour (1.1-gpm) circulating centrifugal pump (at 29.5 feet of head), flow meter with regulating valve, sampling port with three-way valve, a weir compartment for distributing flow in the reaction zone.

Two different light sources were also used for the research. The first was an Ace Glass Incorporated 450-W medium pressure, quartz, mercury-vapor lamp. The second was a Heraeus Noblelight NNI 125/84 XL 150-W bulb, with irradiance at 254 nm of 0.35-mW/cm<sup>2</sup> and radiation flux at 254-nm of 38-W. The irradiation spectrum (Figure 8) shows that the lamp provides most of its intensity from 250 – 260 nm in the UV-C germicidal range. Inside the falling film reactor zone, there is an inner protective tube for the lamp. This tube is made of quartz glass (transmittance = 80-90%) with diameter 43-mm. The reactor wall is made of borosilicate glass with an outside diameter of 110-mm, and the glass tubing is protected with an external tube made of polymethyl methacrylate (PMMA XT) at 140-mm diameter. The borosilicate glass and the PMMA both block the transmittance of UV light at wavelengths less than 300-nm to protect the user.



**Figure 8. Irradiation spectrum for the Strahler NNI 125/84 XL low pressure UV lamp as provided by the manufacturer.**

The UV-C light intensity was measured for both radiation sources using a Sper Scientific 850010 UVC light meter. The results are summarized in Table 8.

**Table 8. Measured UV light intensity.**

Light Source	UV-A&B	UV-C
150-W	0.518 mW/cm <sup>2</sup>	7.21 mW/cm <sup>2</sup>
450-W	56 mW/cm <sup>2</sup>	0.06 mW/cm <sup>2</sup>

### 2.3 FALLING FILM REACTOR

The falling film reactor is the basic setup of how the reactor was originally designed to work for previous experiments conducted by FAU (Meeroff and Youngman 2013). The device is designed to oxidize contaminants in water using UV radiation. The process begins by adding the desired liquid to the 10-L reservoir. The liquid is then pumped up through the flow regulator, which allows a flow range of 30 – 320 Lph. Following the flow regulator, there is a three-way valve which leads to either the sampling port or the weir compartment. The liquid builds up in the weir compartment until it falls over the weir onto the cylindrical reactor wall, which surrounds the UV lamp. While the liquid runs down the reactor wall it is exposed to ultraviolet emittance before it falls back into the reservoir. Underneath the reservoir is a drain valve to remove the desired liquid from the machine.

To operate as a falling film reactor, the leachate is added to the reservoir, and then the light source and the pump is started. Once the leachate has started to circulate, the desired amount of TiO<sub>2</sub> is mixed in to the leachate. The TiO<sub>2</sub> is hydrophobic, so the method used to mix is to add small amount of leachate from the discharge port to a beaker containing the TiO<sub>2</sub>. A slurry is made and then added directly to the reservoir. To ensure all the TiO<sub>2</sub> is added the beaker is rinsed three times with discharge which is then put back into the reservoir.

A flow rate of 300-Lph is maintained by the flow regulator, and the temperature is monitored in the reservoir, during operation.

## 2.4 FLOW THROUGH REACTOR

The flow through reactor is a modification of the falling film reactor design. The process begins by adding the desired amount of leachate to the 10-L reservoir. The liquid is then pumped up through the flow regulator, which allows a flow range of 30 – 320 Lph. Following the flow regulator, there is a three-way valve, which leads to either the sampling port or the weir compartment. The liquid builds up in the weir compartment until it falls over the weir onto the cylindrical reactor wall, and collects in the reactor surrounding the UV lamp in the middle. The leachate slowly drains through the reactor where it is exposed to ultraviolet emittance before passing through the ball valve and discharging back into the reservoir. Underneath the reservoir is a drain valve to remove the leachate from the machine after treatment.

To operate as a flow through reactor, the leachate is added to the reservoir, and then the light source and the pump is started. Once the leachate has started to circulate, the ball valve is partially closed and then the reactor zone fills up. The flow is then regulated to 210-Lph. The desired amount of  $\text{TiO}_2$  is then mixed in to the leachate. The  $\text{TiO}_2$  hydrophobic, so the method used to mix is to add small amount of leachate from the discharge port to a beaker containing the  $\text{TiO}_2$ . A slurry is made and then added directly to the reservoir. To ensure all the  $\text{TiO}_2$  is added, the beaker is rinsed three times with discharge, which is then put back into the reservoir. The flow through reactor can be seen in Figure 9.



**Figure 9: Reactor configured as a flow through.**

In test 4, 20-mL of an antifoam silicone emulsion by J.T. Baker was added at the start of the test to help with foaming. It was added into the reservoir, and almost instantaneous results were seen with the disappearance of foam in the reactor.

## 2.5 IMPROVEMENTS/ADDITIONS

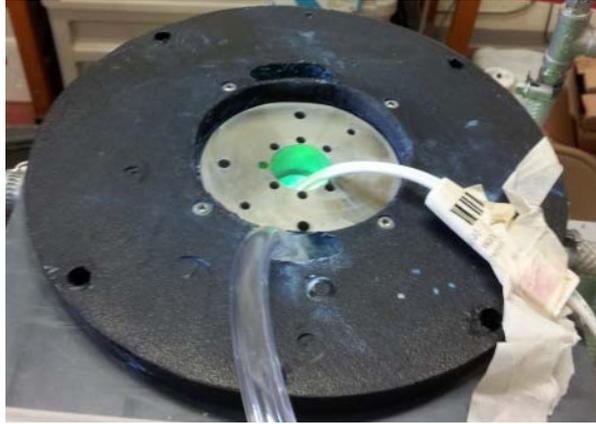
The exposure of leachate to UV light creates an exchange of heat and radiation, causing the leachate to increase in temperature. Tests in the pilot reactor were limited to 4 hours because of the inability to cool the leachate and maintain a constant temperature for kinetics testing. The addition of a 50-foot long, 304-stainless steel cooling coil (Figure 10) in the reservoir attached to a VWR Recirculating Chiller 1150S and filled with 13-L of Dynalene HC-50 (hydro-coolant), enabled extended operation times while maintaining the selected temperature range for the leachate treatment.



**Figure 10. Photograph of the 304 stainless steel cooling coil.**

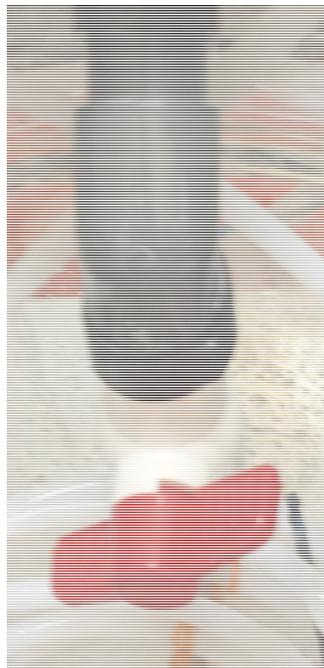
A new stainless steel pump (Speck Pump Y-2951-W-MK) was also fitted to the reactor. This pump, in addition to being less susceptible to corrosion than the previous 260 Lph circulating centrifugal pump (Speck pump model Y-2951.0344 complete with 0.18kW, single-phase motor, 110V, 60Hz, 3600RPM, frame size 56, IP55, C-UL-US, flange turned 90°) has a magnetic coupled drive. This means that no drive shafts collect directly to the impeller, removing the likelihood of leaking from the bearing seals.

Another modification is a new reactor lid created for the flow through test. This new lid has openings so that aeration can be achieved directly in the reactor. The aeration provides mixing to maintain the photocatalytic particles in suspension within the reactor zone and also supplies oxygen for the oxidation process. The lid was custom-made out of HDPE to the exact specifications of the reactor dimensions (Figure 11).



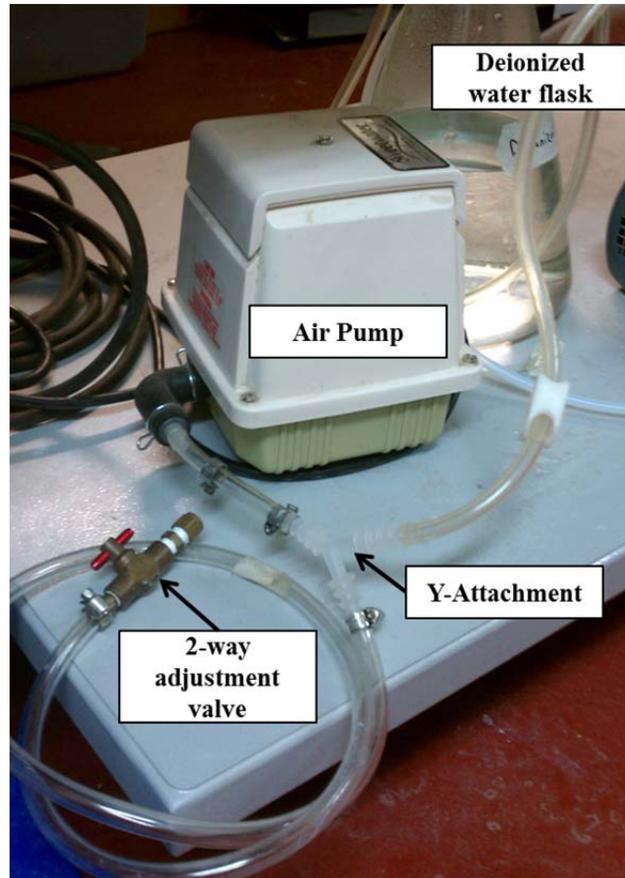
**Figure 11. Photograph of custom flow through reactor lid.**

To create the flow through reactor from the falling fill reactor, a simple 1-inch PVC ball valve was installed in the discharge line of the reactor, see Figure 12. This allows the regulation of the flow from the reactor increasing detention times and allowing us to create a flow through reactor zone.



**Figure 12. Photograph of ball valve.**

An air pump, specifically, the Sweetwater SL22 linear air pump was utilized in conjunction with a large flask of deionized water (to saturate the air with moisture to limit evaporation) is used for aeration. Aeration is needed in the falling film reactor in the reservoir to keep the  $\text{TiO}_2$  in suspension. While in the flow through reactor, aeration is also required to keep the proper dissolved oxygen level for the advanced oxidation process to occur. See Figure 13 for pump system setup.



**Figure 13. Photograph of aeration system.**

## **2.6 EXPERIMENT PROCEDURE**

To initiate the experiment, leachate was removed from the storage refrigerator and 10-L was measured out 2-L at a time using a 2000-mL plastic graduated cylinder (see Figure 14). Once the leachate was added to the reservoir, the unit was powered up and the pump would begin circulating. At this point, a 50-mL sample was taken from the reservoir for testing as the initial concentration at  $t_0$ . Next the  $\text{TiO}_2$  was added in slurry form, the dosing of the  $\text{TiO}_2$  varied from 120-mg/L to 30-g/L. The  $\text{TiO}_2$  was weighed on a Mettler-Toledo XS204 DeltaRange Analytical Balance in a 1000-mL HDPE beaker.



**Figure 14. Measuring leachate in a 2000-mL graduated cylinder.**

Once the  $\text{TiO}_2$  was added, a digital timer was started, and the test began. For the initial runs, 50-mL samples were taken every hour. When the results of the test were examined, it was found that the change over a one hour period was within the range of error, so thereafter, samples were taken at 2-hour increments. Directly after the sample was collected, the temperature, dissolved oxygen content, and pH were taken. The remaining tests were performed after the run was complete.

## **2.7 $\text{TiO}_2$ DOSING**

A widely used, high quality  $\text{TiO}_2$  product (Degussa Aeroxide  $\text{TiO}_2$  P-25) was used as the photocatalyst for all testing. A breakdown of the composition in Aeroxide  $\text{TiO}_2$  P-25 is shown in Table 9.

**Table 9. Physical composition of Degussa Aeroxide TiO<sub>2</sub> P-25.**

Compound	Unit	Value
Titanium Dioxide	wt. %	≥ 99.5
Al <sub>2</sub> O <sub>3</sub>	wt. %	≤ 0.300
SiO <sub>2</sub>	wt. %	≤ 0.200
Fe <sub>2</sub> O <sub>3</sub>	wt. %	≤ 0.010
HCl	wt. %	≤ 0.300
Sieve Residue	wt. %	≤ 0.050

Source: Aeroxide<sup>®</sup> TiO<sub>2</sub> P-25 Product Information Sheet

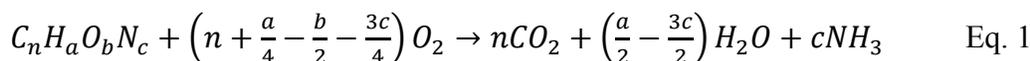
The titanium dioxide in the Degussa Aeroxide TiO<sub>2</sub> P-25 is not a pure form of TiO<sub>2</sub>. Ohtani et al. (2010) tested the crystalline composition of the Aeroxide P-25 and found that it contained a ratio of anatase, rutile and an amorphous phase of the two. They reported that the Degussa P-25 was 78% anatase, 14% rutile and 8% amorphous phase. Some notable chemical and physical properties of the two pure forms as well as the Aeroxide TiO<sub>2</sub> P-25 (used in this study) are listed in Table 10.

**Table 10. Properties of anatase and rutile forms of titanium dioxide.**

Property	Units	Anatase	Rutile	Aeroxide P-25
Molecular Weight	g/mol	79.88	79.88	79.88
Melting Point	°C	1825	1825	1850
Boiling Point	°C	2500-3000	2500-3000	n/a
Light Absorption	nm	<390	<415	<400
Density	g/cm <sup>3</sup>	3.79	4.13	3.8
Crystal Structure	n/a	Tetragonal	Tetragonal	Tetragonal
Refractive Index	n/a	2.55	2.75	2.49
Dielectric Constant	n/a	31	114	78.5

Source: Pelaez et al. (2012), Hong et al. (2005), Faure et al. (2010), Kosmulski et al. (2009), Evonik Industries (2008)

Finding the proper dose of TiO<sub>2</sub> is vital to having the advanced oxidation process work correctly. The conventional method would be to have a balanced molar equation for the substance to remove. However, since leachate is a cocktail of different contaminants, to determine the exact value using stoichiometry is complicated. When considering a basic molecule of an organic compound containing carbon, hydrogen, oxygen, and nitrogen to be oxidized to carbon dioxide, ammonia and water, the following chemical reaction is can be used:



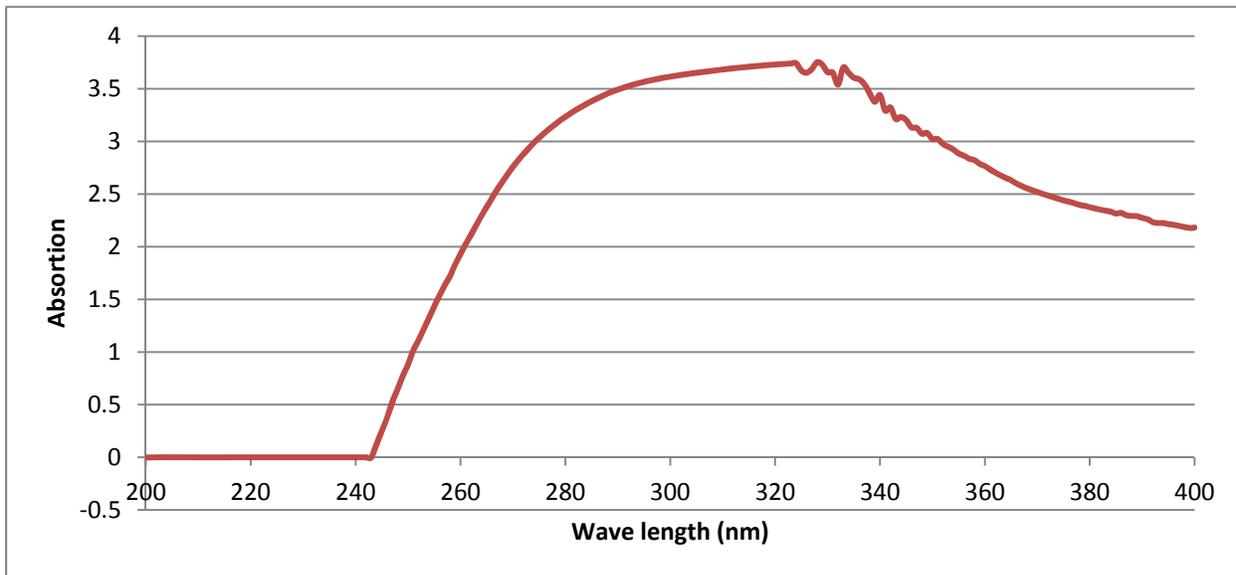
Assuming the average organic compound found in leachate is derived from a mixture of alkaloids, lipids, proteins and peptides, this theoretical molecule would contain six carbon atoms. From the chemical formula in Equation 1, then  $n = 6$ . Since the ratio is greater than 6:1 for protons from the TiO<sub>2</sub>, then it would be expected that the theoretical TiO<sub>2</sub> dose should be 6 times

the COD value. It is important to note, that the theoretical dose may limit the distance that the UV light can travel in the bulk leachate solution. In other words, too much  $\text{TiO}_2$  will cause only surface radiation, while too little  $\text{TiO}_2$  will cause UV light to pass directly through the solution without activating the catalyst.

To determine the level of UV light absorbance, a full scan was performed on a DR5000 Spectrometer from 200-nm to 400-nm, the UV range. The cuvettes are 1-inch in diameter, which is slightly smaller than the inner diameter of the reactor. The peak absorbance was found to occur at 330-nm as shown in Figure 16. Dilutions of  $\text{TiO}_2$  and deionized water were made from 30-g/L to 0.0624-g/L and tested.



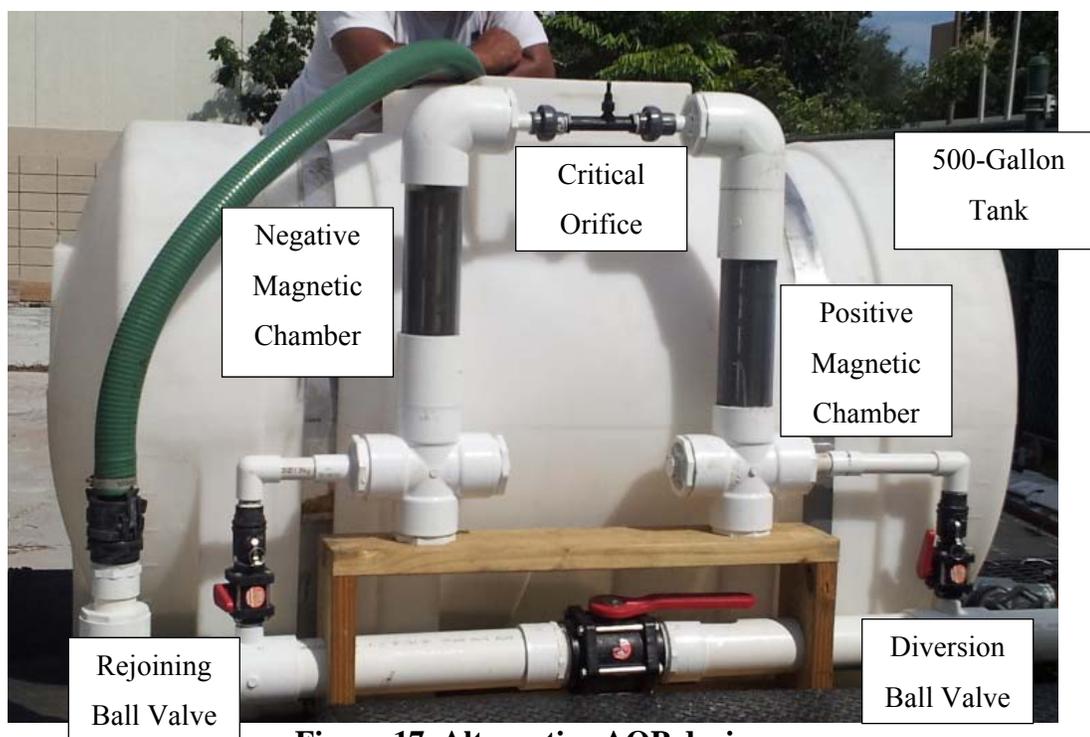
**Figure 15: Dilutions created for a standard Beer's Law curve.**



**Figure 16: Full spectrum absorption curve from 200-400-nm using the DR4000U.**

## 2.8 CRITICAL ORIFICE ADVANCED OXIDATION PROCESS

In this research, an alternative advanced oxidation process was also tested. The set-up is a simple one for the pilot scale. It consists of a 500-gallon tank, a 150-gpm Honda pump, a positive magnetic chamber, a negative magnetic chamber, and a critical orifice (see Figure 17 for labeled components).



**Figure 17. Alternative AOP device.**

10-gallons of leachate was collected from Dyer Park and poured in to the 500-gallon tank, it was diluted with 50-gallons of tap water. The pump was turned on, and once the flow was started, a 50-mL sample was taken. Then the rejoining ball valve and the diversion valves were opened causing the leachate to flow through the reaction chamber. Leachate flows from the tank to the pump, and then the pump sends the leachate to the diversion valve. Only about 10% of the leachate is sent through the reactor, as controlled by the diversion ball valve. The rest (~90%) is diverted around the reactor. The leachate that is diverted passes through a positive magnetic field created by neodymium magnets and stainless steel rods. Then the leachate is passed through a critical orifice, where it is pressurized and then ejected at a high velocity creating a vacuum where the dissolved oxygen comes out of solution. This oxygen creates thousands of micro-bubbles with a relatively large combined surface area. This oxygen oxidizes the COD, then the micro-bubbles and leachate travel through the negative chamber and are reintroduced to the by-pass stream of leachate where the micro-bubbles continue to oxidize COD. Samples were taken at 5-minute intervals from a discharge port.

## 2.9 CRYSTAL VIOLET TEST

The crystal violet test was used to verify the generation of hydroxide ion in the reaction mechanism. The crystal violet test was selected because it is a cation and in the presence of the hydroxide ion, it will be neutralized. This neutralization is visually observed as the discoloration of the dye. This test was performed in a photocatalytic chamber using the 450-W lamp. A glass petri dish containing 25-mL of leachate with 10-g/L  $\text{TiO}_2$  was prepared. Added to this was 5-mL of  $1.0 \times 10^{-4}$  M crystal violet. The lamp was turned on in the empty chamber and allowed to warmup for 15-minutes. Then the sample was placed in the chamber at 4-inches from the light, and observations were taken at 1-minute intervals.

## 2.10 IMPROVING COD REMOVAL

To improve the COD removal process, different catalysts were tried in the photocatalytic test chamber. A stock solution of 5-g/L of  $\text{TiO}_2$  and leachate was made. Then 80-mL of the solution was added to quartz test tubes. Then different catalysts were added to eleven test tubes. The catalysts tested included: aluminum, zinc, steel wool, and combinations of these catalysts. The 450-W light was warmed up in the empty chamber for 15-minutes. Then the test tubes were added at a distance of 4-inches, see Figure 18. The test was run for one hour after which 2-mL samples were taken from each test tube and tested for COD. The remainder of the leachate was properly disposed.

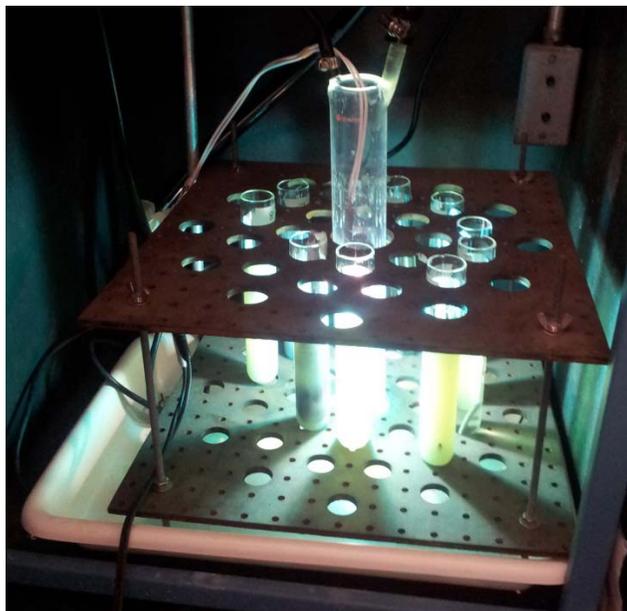


Figure 18. Improved COD removal test chamber setup.

## 2.11 ANALYTICAL METHODS FOR PARAMETERS OF INTEREST

The advanced oxidation unit was used to test the removal efficiency of the following constituents: COD, ammonia, alkalinity, pH, and temperature.

### 2.11.1 COD

For chemical oxygen demand (COD) testing the Reactor Digestion Method for the Hach DR5000U was used with the High Range COD digestion vials (20 to 1,500 mg/L as  $\text{O}_2$ ). The method relies on the reduction of the orange dichromate ion ( $\text{Cr}_2\text{O}_7^{2-}$ ) to the green chromic ion ( $\text{Cr}^{3+}$ ), which is analyzed colorimetrically. Theoretically, since each dichromate ion accepts 6 electrons per molecule and each molecule of dioxygen accepts 4 electrons, the COD of 1-g of  $\text{Cr}_2\text{O}_7^{2-}$  is equal to 1.5-g of molecular oxygen. The COD was tested prior to treatment and then after every subsequent 2-4 hours of treatment. At least 2 duplicate samples were created for each COD test. After centrifugation at 6000 rpm for 8 minutes to separate the photocatalyst particles, samples were diluted to 2.0-mL using 18.2M $\Omega$ -cm deionized water. Throughout the testing, a dilution of 1:2 was used. Briefly, 2.0-mL samples were inverted 20 times to mix and placed in a heating block at 150°C to digest for 2-hours. Samples were removed from the heating block and

inverted another 20 times before being allowed to cool for one hour in the dark. At this point, samples were wiped clean using a Kim-Wipe and analyzed using a Hach DR5000U UV-Vis spectrophotometer. The COD value in mg/L was recorded. A certified reference material (Total Organic Carbon Std. 1,000ppm prepared to EPA Method 415-1 Aqua Solutions, Deer Park, TX) was used to check the instrument calibration. This value varied from 0.1-10% error. One calibration check standard was analyzed per batch of samples. It was discovered that high levels of chloride in the samples can interfere in this test method. When accelerated levels of chloride (>2,000 mg/L) are present in the sample, the chlorides can be quantitatively oxidized by the dichromate, consequently displaying erroneous levels of oxidizable organic compounds. Due to sample dilution, high levels of chlorides that would affect the COD test were not encountered.

### 2.11.2 Ammonia-Nitrogen

For ammonia-nitrogen, the EPA Method # 350.2 (Detection of Ammonia by Colorimetry), Nessler spectrophotometric method was used. An ammonia (Medium Range: 0.00 to 9.99 mg/L as  $\text{NH}_3\text{-N}$ ) ion specific meter from Hanna Instruments (Woonsocket, RI) was used. The test began by adding 10-mL of sample to a 10-mL cuvette. The outside of the cuvette was cleaned with a Kim-wipe to remove any fingerprints or dust and placed in the instrument to be zeroed out. Next, four drops of the first reagent (HI 93715A) which is a mineral stabilizer and polyvinyl alcohol dispersing agent were added to the cuvette and the solution was swirled. Then four drops of the second reagent (HI 93715B) which is Nessler's reagent were added, and the cuvette was swirled (see Figure 19) again.



**Figure 19. Reagent addition to cuvette for ammonia testing.**

The outside of the cuvette was cleaned once more before being placed back into the instrument. The recommended reaction time of 3.5-minutes was allowed to pass before the reading was taken. The instrument directly displayed the concentration in mg/L of ammonia-nitrogen ( $\text{NH}_3\text{-}$

N) on the liquid crystal display (see Figure 20).



**Figure 20. Sample reading from the HI 95715 Ammonia Medium Range ISM.**

To convert the reading to mg/L of ammonia ( $\text{NH}_3$ ), ammonia-nitrogen values can be multiplied by a factor of 1.216. Dilution was necessary to lower the concentrations of known interferences, such as organics, sulfides, color, chloramines, aldehydes, and hardness to below 1-g/L as  $\text{CaCO}_3$ . Due to the potency of the ammonia, a 20% dilution was used.

### **2.11.3 Alkalinity**

For the total alkalinity measurements, SM 2320B was used. A Hach digital titrator was loaded with a 1.600-N  $\text{H}_2\text{SO}_4$  titrant cartridge for all analyses. First the sample was diluted with sterile deionized water (range: undiluted to 1:10), then phenolphthalein indicator was added to the sample. Titrant was added until the phenolphthalein endpoint was reached (pink to clear), if necessary. The reading on the digital titrator was recorded as corresponding to the phenolphthalein alkalinity in mg/L as  $\text{CaCO}_3$  by multiplying the dilution factor by the number of digits. No phenolphthalein alkalinity was measured during any of the experiments. Then the bromocresol green-methyl red indicator was added to the sample, and again titrant was added until the second endpoint was reached (blue-green to light pink). This is the bromocresol green methyl-red alkalinity. When the phenolphthalein alkalinity and bromocresol green methyl-red alkalinity values were added together, this corresponded to the total alkalinity value.

### **2.11.4 pH**

For all experiments, pH was recorded during, prior, and at the end of every experiment using pH Indicator Strips (Whatman Inc., Clifton, NJ), a Hach SensIon 3 pH meter, a Hach MP-6

multiparameter unit, or a Hach HQ40d Portable pH, Conductivity, Dissolved Oxygen (DO), and ORP probes with a Hach ISE Multi-Parameter Meter, with the latter being used for nearly all of the measurements reported. Probes were calibrated periodically with standard pH buffers (4, 7, and 10). Sensors were rinsed with deionized water and dried with kimwipes in between sample readings.

### 2.11.5 Dissolved Oxygen

After the first two runs of the pilot reactor, the dissolved oxygen was measured. The dissolved oxygen needs to be known to insure that enough oxygen is available to react with the  $\text{TiO}_2$  to remove metals. A Hach HQ40d Portable pH, Conductivity, Dissolved Oxygen (DO), and ORP probe was used to measure the dissolved oxygen content in the samples as soon as they were removed from the reactor.

### 2.11.5 Temperature

The temperature was recorded prior to and during all experiments. The advanced oxidation machine has a built in temperature probe in the 10-L reservoir. The temperature was recorded from the digital output located on the control box of the unit. The temperature was recorded every hour. The leachate temperature was also taken during the DO reading for QA/QC of reservoir thermometer.

## 2.12 CATALYST RECOVERY

An attempt to collect the used catalyst from the treated leachate for reuse was tried two different ways. The first was sedimentation of a sample of leachate that was taken from the reactor in a glass beaker. The beaker was then allowed to settle quiescently. Pictures were taken at 1-minute increments, see Figure 21.



Figure 21. Sedimentation time lapse series over a 10-minute period.

The second trial was filtering with a 1-micron bag filter. The leachate was allowed to settle in a 5-gallon bucket. The dry filter bag was weighed. Then the filter bag was placed over a second 5-gallon bucket, and the leachate was poured through the filter.

### 2.13 INTENSITY OF UV LIGHT

To measure the amount of UV light that the  $\text{TiO}_2$  is exposed to in the falling film and the flow through reactor the UV intensity was measured. The UV light was measured in three spectrums the: UVA, UVB and UVC ranges using a Fisher Scientific UV light meter 06-662-65 for UVA and UVB and a Sper Scientific 850010 UVC light meter.

The lamps were allowed to warm up for 15-minutes to the correct operating temperature. Then the measuring devices were placed 0.5-inches from the light, and a measurement was taken. This was repeated three times for each light source, and an average of the readings was taken. To calculate the intensity, the arc length of the light and the outer circumference of the lamps protective tube were multiplied to get the total exposure area. This was then multiplied with the measured UV intensity for each spectrum, giving the total power of the lamp to that surface. This was then multiplied by the number of times the leachate volume is recirculated per hour and then by the number of hours the test was conducted for to find the Watt-hours per liter of exposure.

$$\text{Light Intensity} \left( \frac{W \times hr}{L} \right) = \text{Exposure Area} (cm^2) \times \text{Measured Intensity} \left( \frac{W}{cm^2} \right) \times$$

$$\text{Retention time} \left( \frac{Pass}{L \times hr} \right) \times \text{Exposure Time} (hr) \times \text{Length of Test} (hr)$$

### **3. RESULTS AND DISCUSSION**

This chapter explains in detail the results obtained from the experimental procedures that were described in the methodology section.

#### **3.1 FALLING FILM EXPERIMENT #1**

The falling film experiment was conducted with both the 150-W and the 450-W lamps. Each was run for 8-hours, with samples taken at 1-hour intervals. For the first set of trials, the TiO<sub>2</sub> concentration was set at 5-g/L. The COD removal for the two trials was 0% for the 150-W lamp and 17% for the 450-W lamp. The ammonia removal was 7% for the 150-W lamp and 3% for the 450-W lamp. The alkalinity removal for the 150-W lamp was 9%, while the 450-W lamp was 2%. Results can be seen in Table 11 and Table 12.

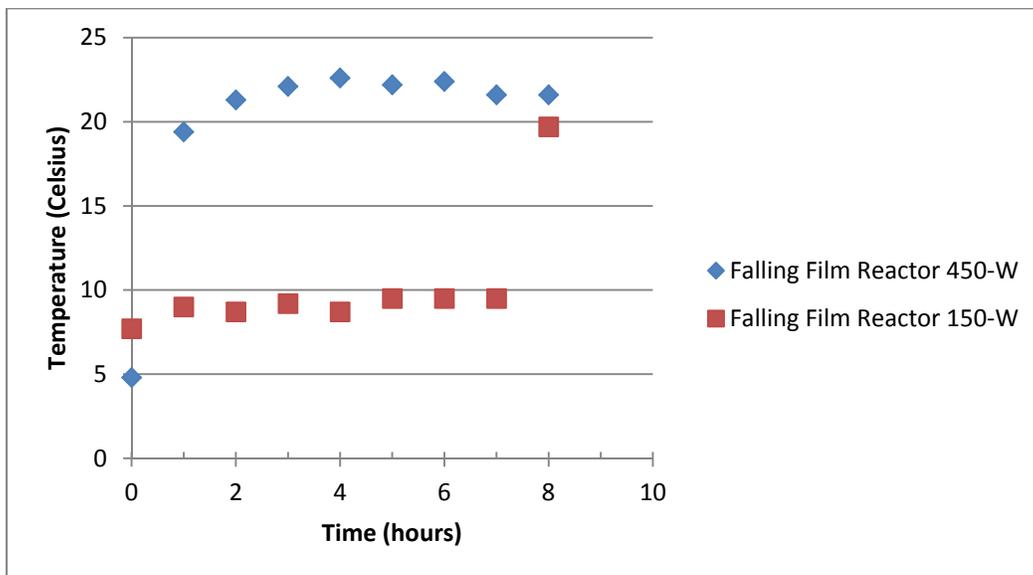
**Table 11. Falling film reactor experiment using 150-W light source (TiO<sub>2</sub> = 5-g/L).**

Sample	Temp. °C	KWH	COD mg/L	COD % remaining	NH <sub>3</sub> mg/L	NH <sub>3</sub> % remaining	pH	Alkalinity mg/L as CaCO <sub>3</sub>	Alkalinity % remaining
0	7.7	--	375	--	313	--	7.35	1550	--
0	7.7	0.52	300	100.0	248	100.0	8.05	1120	100.0
1	9.0	1.74	270	90.0	237	95.6	8.11	1100	98.2
2	8.7	2.90	275	91.7	244	98.4	8.15	1110	99.1
3	9.2	4.17	290	96.7	248	100.0	8.06	1050	93.8
4	8.7	5.61	275	91.7	226	91.1	8.04	1080	96.4
5	9.5	6.72	265	88.3	253	102.0	7.98	1050	93.8
6	9.5	6.72	265	88.3	254	102.4	7.98	1050	93.8
7	9.5	7.64	270	90.0	253	102.0	8.08	1040	92.9
8	19.7	8.56	300	100.0	232	93.5	8.20	1020	91.1

**Table 12. Falling film reactor experiment using 450-W light source (TiO<sub>2</sub> = 5-g/L).**

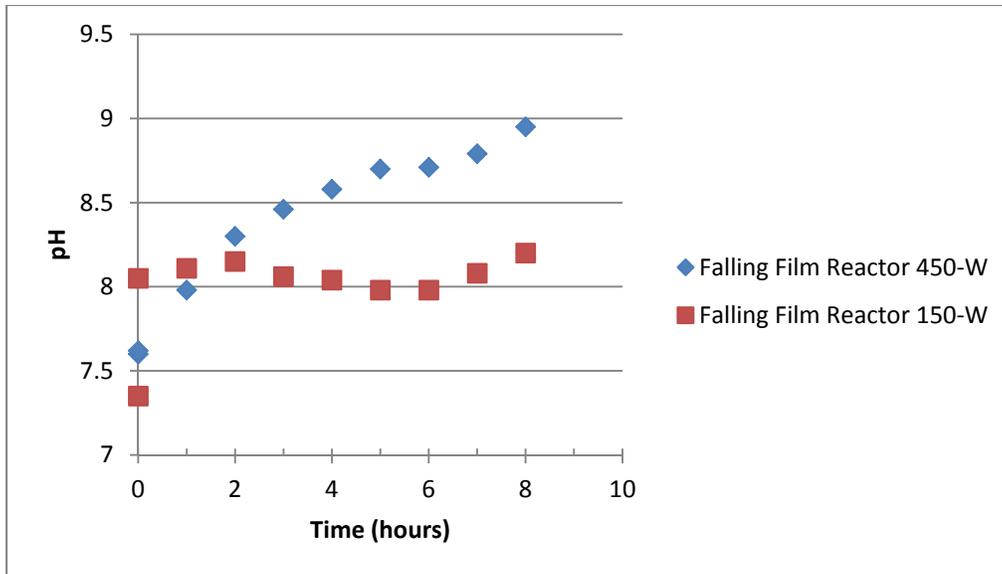
Sample	Temp. °C	KWH	COD mg/L	COD % remaining	NH <sub>3</sub> mg/L	NH <sub>3</sub> % remaining	pH	Alkalinity mg/L as CaCO <sub>3</sub>	Alkalinity % remaining
0	4.8	--	375	--	313	--	7.60	1550	--
0	4.8	0	305	100.0	250	100.0	7.62	1070	100.0
1	19.4	1.39	290	95.1	248	99.0	7.98	1090	101.9
2	21.3	2.77	295	96.7	220	87.9	8.30	1090	101.9
3	22.1	4.16	280	91.8	227	90.8	8.46	1050	98.1
4	22.6	5.54	290	95.1	232	92.7	8.58	1040	97.2
5	22.2	6.93	280	91.8	233	93.2	8.70	1030	96.3
6	22.4	8.31	280	91.8	218	86.9	8.71	1020	95.3
7	21.6	9.70	275	90.2	230	91.7	8.79	1000	93.5
8	21.6	11.08	255	83.6	243	97.1	8.95	1050	98.1

The falling film temperature trend (Figure 22) shows that the 450-W lamp transfers more energy into the leachate than the 150-W lamp. There is a clear difference in the initial temperatures of the two tests, cooling was reduced at the seven hour mark of the 150-W test to see if cooling could be kept consistent between the two tests. The results show that the new cooling system is capable of cooling the leachate for continuous running of the experiment. However, the proper temperature set point has not been determined yet. The most economical and ideal set point for a full scale plant would be to provide no cooling at all; however, constant temperatures are required for kinetics testing for eventual scale up.



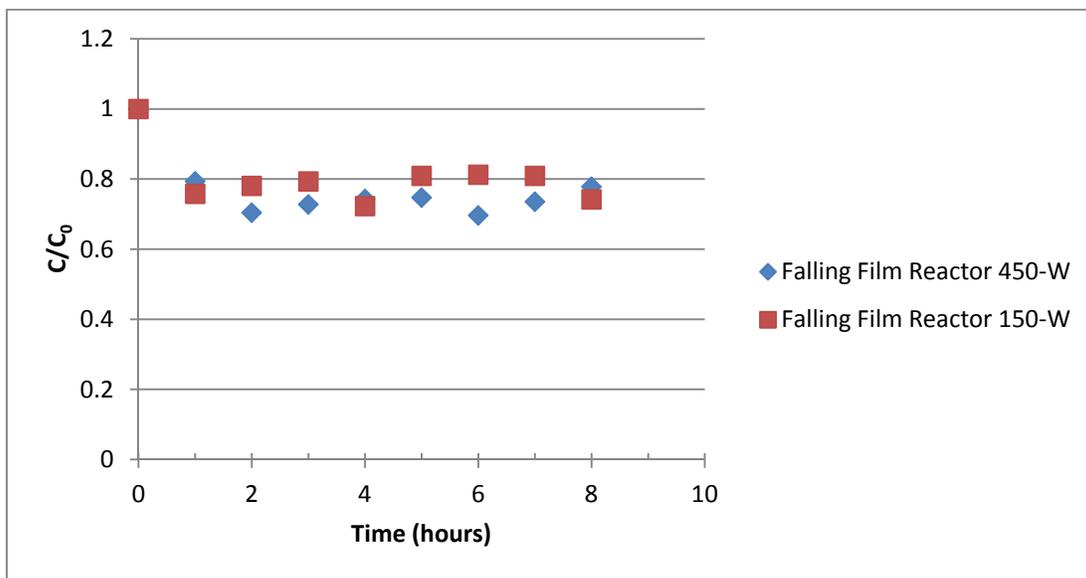
**Figure 22. Falling Film Reactor Temperature Comparison.**

The falling film pH trend (Figure 23) shows that the 450-W lamp had a more rapid increase in the pH when compared to the 150-W lamp. The 450-W lamp with a higher pH may have an effect on the treatment efficiency, but this has not been quantified yet at this time. Both tests show that the TiO<sub>2</sub> and exposure to UV light increase the pH. The hypothesis is that as the photocatalytic action of the TiO<sub>2</sub> increases the pH, the decomposition of organics should produce CO<sub>2</sub> degassing, which should alter the pH and buffering capacity of the bulk solution.



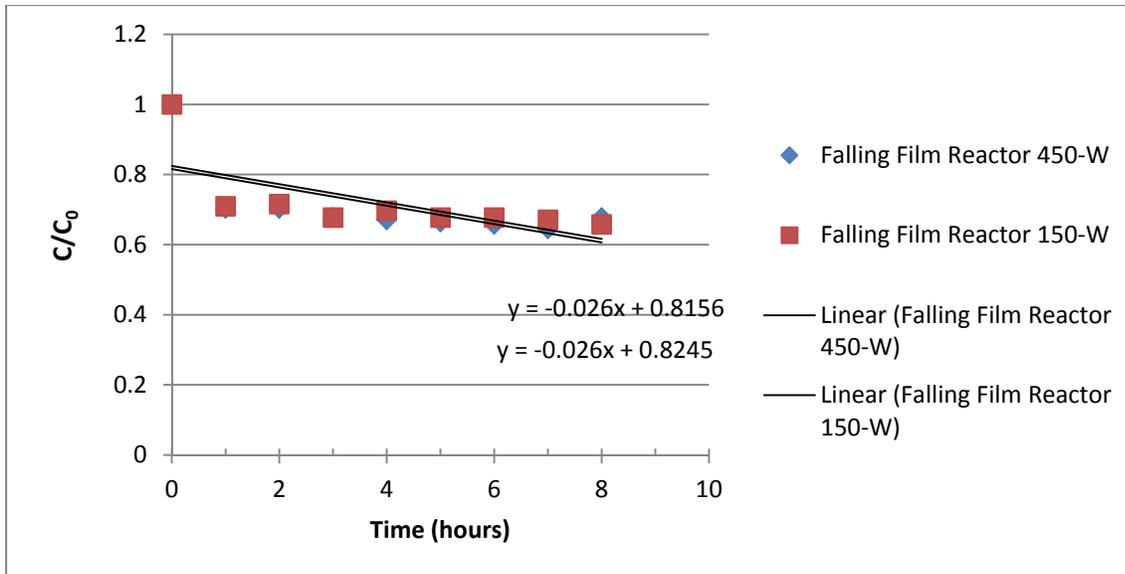
**Figure 23. Falling Film Reactor pH Comparison.**

The falling film reactor ammonia experiment (Figure 24) showed some small reduction in ammonia for both lamps tested. The variability in the ammonia values over time was attributed to the high level of dilution at 1:100 that was used for this test. Small errors in pipetting could lead to large error in values. This was corrected for all following experiments.



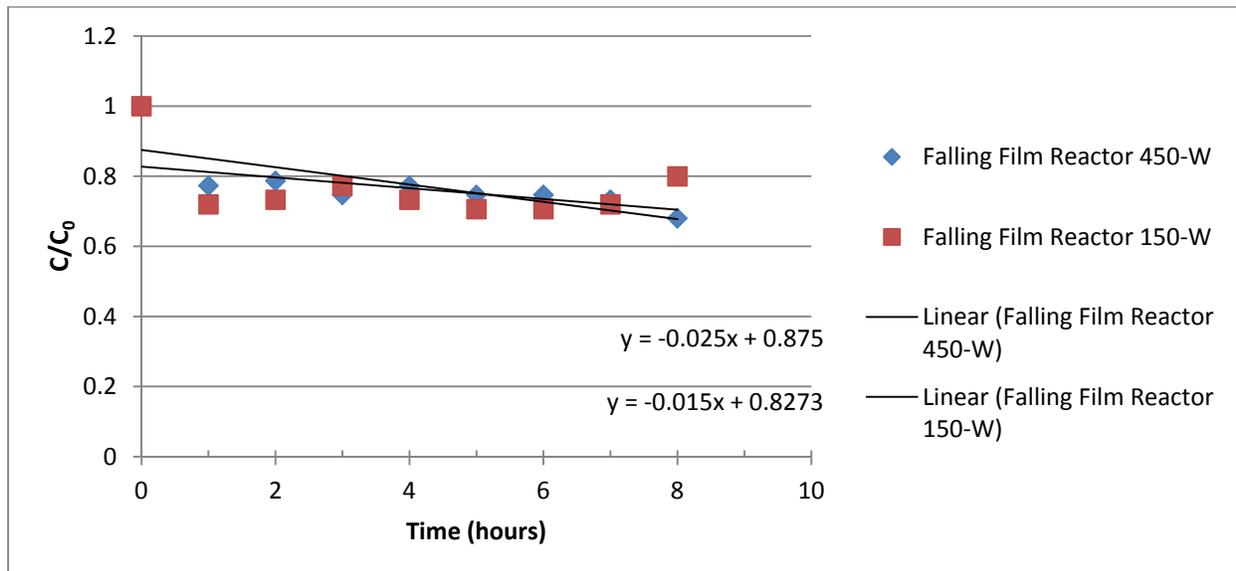
**Figure 24. Falling Film Reactor Ammonia Comparison.**

The falling film reactor alkalinity experiment (Figure 25) showed that both lamps reduced alkalinity at similar rates ( $m_{450-W} = -0.026/hr$  and  $m_{150-W} = -0.026/hr$ ). The removal rate is linear and of the first order. However, the results show that similar to the ammonia readings, the high level of dilutions 1:100 leads to an exaggerated error. The dilution procedures were corrected for subsequent testing.



**Figure 25. Falling Film Reactor Alkalinity Comparison.**

The falling film COD experiment (Figure 26) shows that COD removal is taking place, and that each lamp is removing COD at approximately the same rate of first order removal ( $m_{450-W} = -0.025/\text{hr}$  and  $m_{150-W} = -0.015/\text{hr}$ ). Large spikes at the end of the test were seen for the both the 150-W and 450-W lamps. However, the spikes were in different directions. The 450-W increased removal while the 150-W showed a decrease in removal. Again this is attributed to the dilution that was used, which may have caused an exaggerated error.



**Figure 26. Falling Film Reactor COD Comparison.**

Although this was the first run of the falling film reactor, many operational issues were addressed after this experiment. The best COD removal observed was 17% by the 450-W lamp. There was no apparent difference in the performance of the lamps with respect to removal of the parameters of interest other than the pH levels. The pH of the 450-W lamp rose to 8.95 while the 150-W lamp only went up to 8.20, which was unexpected since the alkalinity levels were essentially the same. There was a clear difference in the temperature of the leachate during test, with the 450-W lamp having an operating temperature of 21.5°C while the 150-W lamp only had a temperature of 9.5°C. Typically, doubling the temperature like this can more than double the reaction rate. Using the Arrhenius temperature relationship, the rate would be expected to be 4.5 times faster for the 450-W lamp in this case.

$$k_{T1} = k_{T2}\theta^{(T1-T2)}$$

**$\theta = 1.135$  if T is between 4 - 20°C**

The reason for the temperature difference can be directly related to the lamp wattage, since the higher watt lamp increased the temperature more rapidly.

### **3.2 FLOW THROUGH EXPERIMENT #2**

The flow through test was conducted with both the 150-W and the 450-W lamps. Each was run for 8-hours, with samples taken at 1-hour intervals using a TiO<sub>2</sub> concentration of 10-g/L. The COD removal for the two trials was 25% for the 150-W lamp and 8% for the 450-W lamp. The ammonia removal was 17% for the 150-W lamp and 46% for the 450-W lamp. The alkalinity removal for the 150-W lamp was 12%, while the 450-W lamp was 24%. Results can be seen in Table 13 and Table 14.

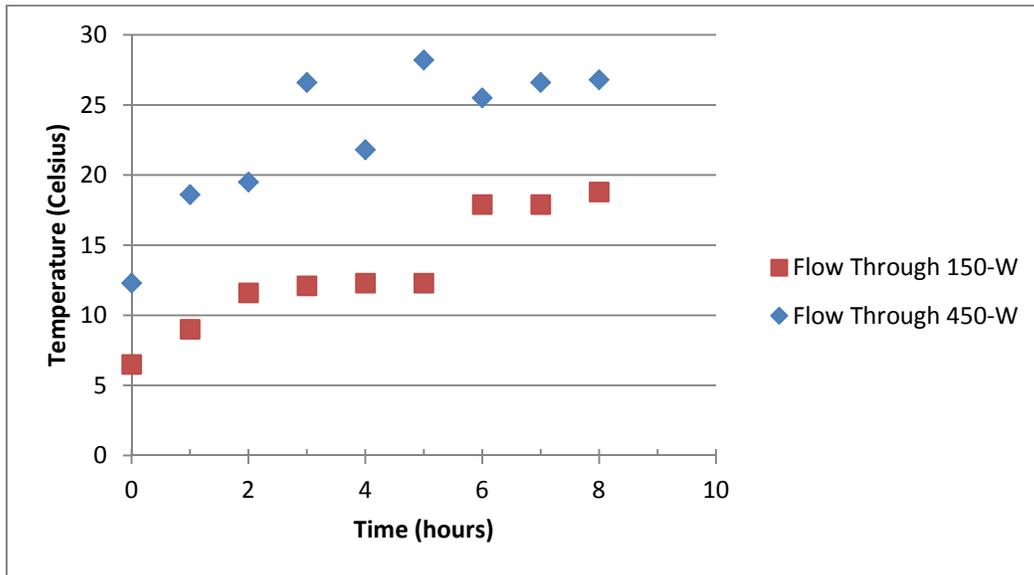
**Table 13. Flow through reactor experiment using 150-W light source (TiO<sub>2</sub> = 10-g/L).**

Sample	Temp. °C	KWH	COD mg/L	COD % remaining	NH <sub>3</sub> mg/L	NH <sub>3</sub> % remaining	pH	DO mg/L	Alkalinity mg/L as CaCO <sub>3</sub>	Alkalinity % remaining
0	6.5	0.52	300	100.0	232	100.0	8.20	9.37	1020	100.0
1	9.0	1.74	285	95.0	215	92.7	8.56	8.67	1050	102.9
2	11.6	2.90	335	111.7	169	72.8	8.48	8.85	1070	104.9
3	12.1	4.17	225	75.0	170	73.3	8.58	8.90	980	96.1
4	12.3	5.61	260	86.7	178	76.7	8.54	8.81	990	97.1
5	12.3	6.72	245	81.7	177	76.3	8.59	8.91	1010	99.0
6	17.9	--	285	95.0	189	81.5	8.60	9.24	880	86.3
7	17.9	7.64	350	116.7	163	70.3	8.73	9.20	970	95.1
8	18.8	8.56	225	75.0	193	83.2	8.7	9.14	900	88.2

**Table 14. Flow through reactor experiment using 450-W light source (TiO<sub>2</sub> = 10-g/L).**

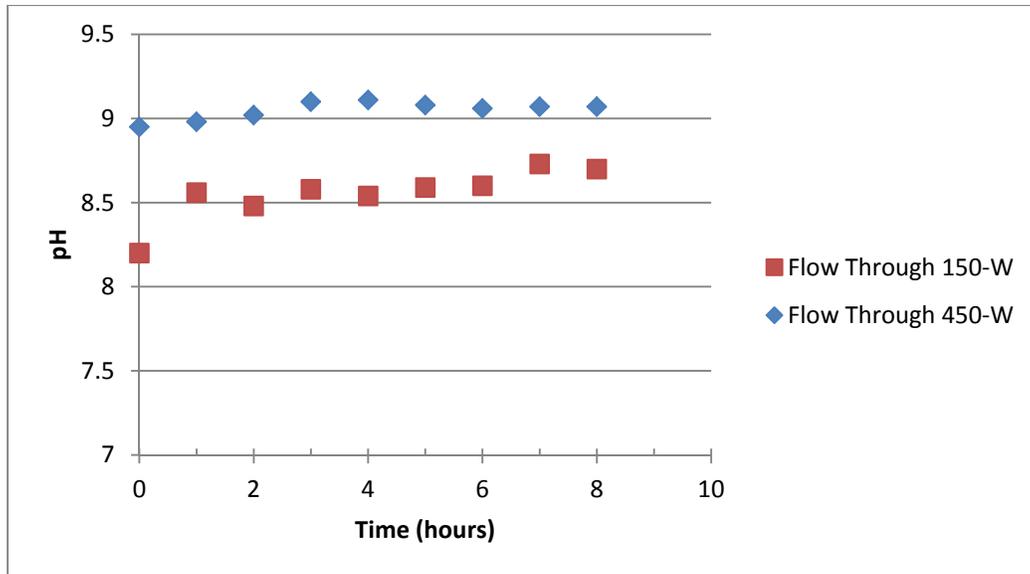
Sample	Temp. °C	KWH	COD mg/L	COD % remaining	NH <sub>3</sub> mg/L	NH <sub>3</sub> % remaining	pH	DO mg/L	Alkalinity mg/L as CaCO <sub>3</sub>	Alkalinity % remaining
0	12.3	0.52	255	100.0	243	100.0	8.95	9.05	1050	100.0
1	18.6	1.74	240	94.1	221	90.8	8.98	9.07	1030	98.1
2	19.5	2.90	235	92.2	206	84.6	9.02	9.02	950	90.5
3	26.6	4.17	235	92.2	202	83.0	9.10	7.34	940	89.5
4	21.8	5.61	250	98.0	172	70.6	9.11	8.7	800	76.2
5	28.2	6.72	245	96.1	161	66.4	9.08	7.48	790	75.2
6	25.5	--	240	94.1	151	62.3	9.06	8.32	820	78.1
7	26.6	7.64	245	96.1	144	59.1	9.07	7.79	770	73.3
8	26.8	8.56	235	92.2	133	54.7	9.07	8.00	700	76.2

The flow through reactor experiment temperature trend (Figure 27) shows that when the liquid is retained near the light for longer periods of time, the temperature increases more rapidly compared to the falling film reactor configuration. Both lamps showed an asymptotic rise in temperature that stabilized near the end of the run at 26.8°C for the 450-W lamp compared to the falling film temperature of 21.5°C. The amount of time the leachate spent in the reservoir is much lower than the falling film test. The flow through reactor detention time is approximately 42 minutes per hour, while the falling film reactor detention time is approximately one minute per hour. Similar results are seen for the 150-W lamp with the final flow through temperature at 18.8°C compared to the falling film reactor configuration at 9.5°C. According to the Arrhenius relationship, this would account for a difference of 2.75 times in the reaction rate.



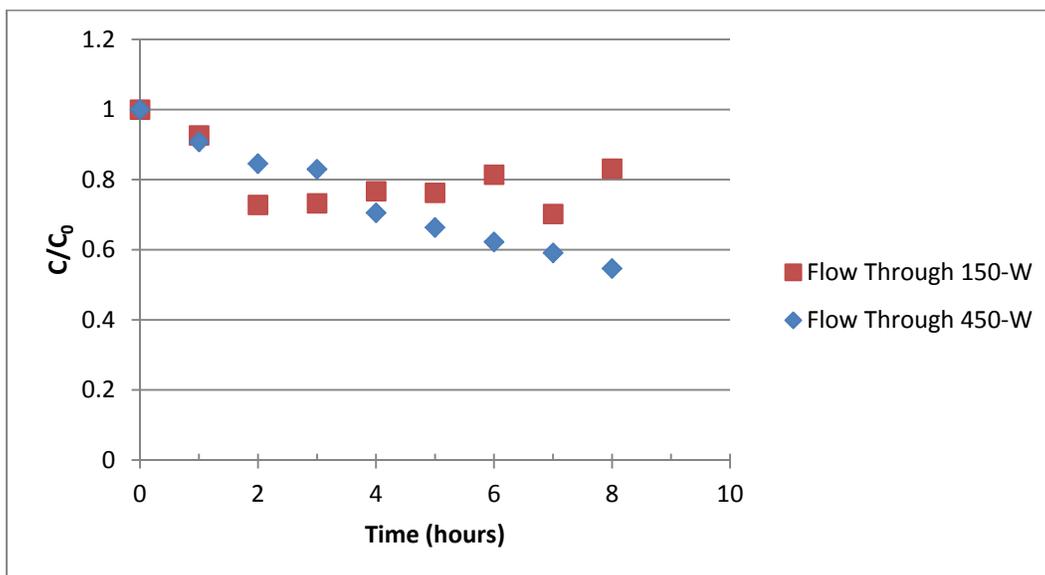
**Figure 27. Flow Through Reactor Temperature Comparison.**

The flow through reactor experiment pH trend (Figure 28) started higher than the falling film test because the leachate was the same that was used in the falling film test. The pH remained in the range of 9.0 for the 450-W lamp. For the 150-W lamp the pH remained in the range of 8.5. Both lamps showed no appreciable decrease in pH during the 8-hours of testing suggesting that the remaining alkalinity buffered any pH changes over the duration of the experiment.



**Figure 28. Flow Through Reactor pH Comparison.**

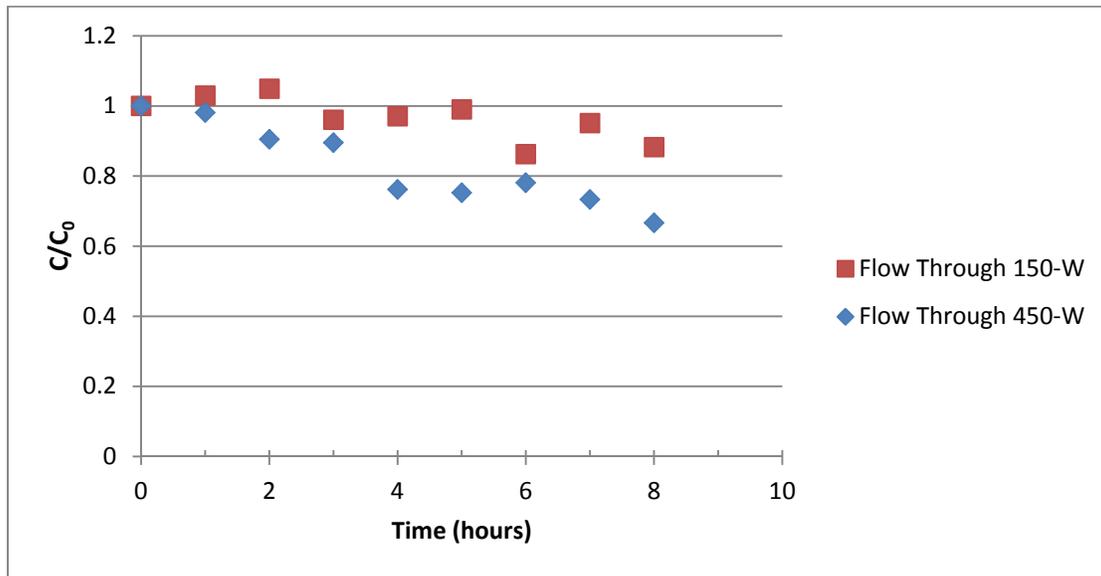
The flow through reactor experiment ammonia trend (Figure 29) showed a 45.3% reduction in ammonia for the 450-W lamp, compared to the 2.9% reduction from the previous 450-W falling film test. The 150-W flow through reactor configuration showed an ammonia reduction of 26.8% compared with the previous 150-W falling film test of 6.5%. Both lamps showed a remarkable improvement in ammonia removal while in the flow through reactor configuration.



**Figure 29. Flow Through Reactor Ammonia Comparison.**

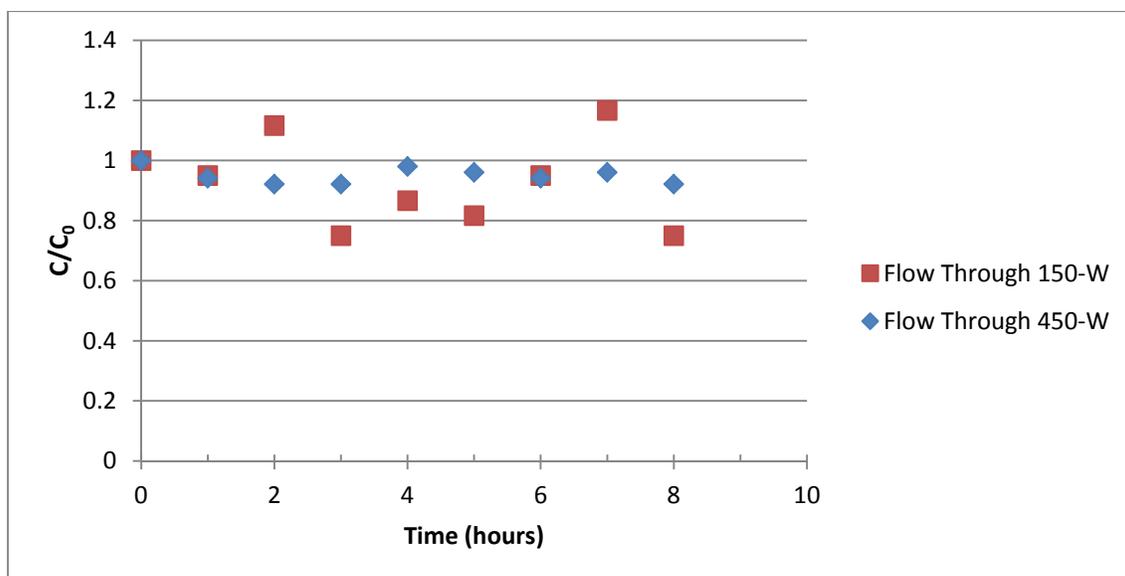
The flow through reactor experiment alkalinity trend (Figure 30) shows that the 450-W light removed 45.3% of the ammonia compared to the 450-W falling film reactor configuration, which

removed only 2.9%. While the 150-W light in the flow through reactor experiment removed 16.8% compared to the previous flow through reactor configuration ammonia removal of only 6.5%. Both lamps demonstrated that the flow through reactor configuration was more effective at ammonia removal; however, this may also be attributed to the higher temperatures of the test.



**Figure 30. Flow Through Reactor Alkalinity Comparison.**

The flow through reactor experiment COD removal trend (Figure 31) showed only a small amount of COD removal from both the 450-W and the 150-W lamps. The 450-W light removed 7.8% compared with the falling film reactor configuration COD removal of 16.4%. The flow through reactor for the 150-W lamp showed 25% removal after 8 hours, while the falling film reactor showed essentially no removal over the same time period. From previous work (Meeroff and Youngman 2014), the COD removal is expected to have the lowest removal rate of the parameters tested.



**Figure 31. Flow Through Reactor COD Comparison.**

This was the first run of the reactor configured as a flow through reactor. It should be noted that the starting COD of the 450-W lamp was much lower to start with compared to the falling film reactor tests, but the ending COD values are close with 225 mg/L for the 150-W lamp and 235 mg/L for the 450-W lamp. The COD was lower because leachate from previous tests was used in this experiment. There was a disparity in the removal of ammonia from the 150-W and the 450-W lamps with 17% removal and 46% removal respectively. This may be attributed to the higher leachate temperature in the 450-W test. The removal of COD is not below the 120 mg/L goal (EPA 2009) nor is the alkalinity below the 20 mg/L as CaCO<sub>3</sub> goal. But improvements were made in removal percentages relative to the 5-g/L TiO<sub>2</sub> falling film reactor experiment #1.

### **3.3 FLOW THROUGH WITH AERATION EXPERIMENT #3**

The flow through reactor experiment with aeration was conducted with both the 150-W and the 450-W lamps. Each was run for 8-hours, with samples taken at 2-hour intervals. The TiO<sub>2</sub> concentration was 10-g/L. The COD removal for the two trials was 8% for the 150-W lamp and 12% for the 450-W lamp. The ammonia removal was 39% for the 150-W lamp and 25% for the 450-W lamp. The alkalinity removal for the 150-W lamp was 32%, while the 450-W lamp was 0%. Results can be seen in Table 15 and Table 16.

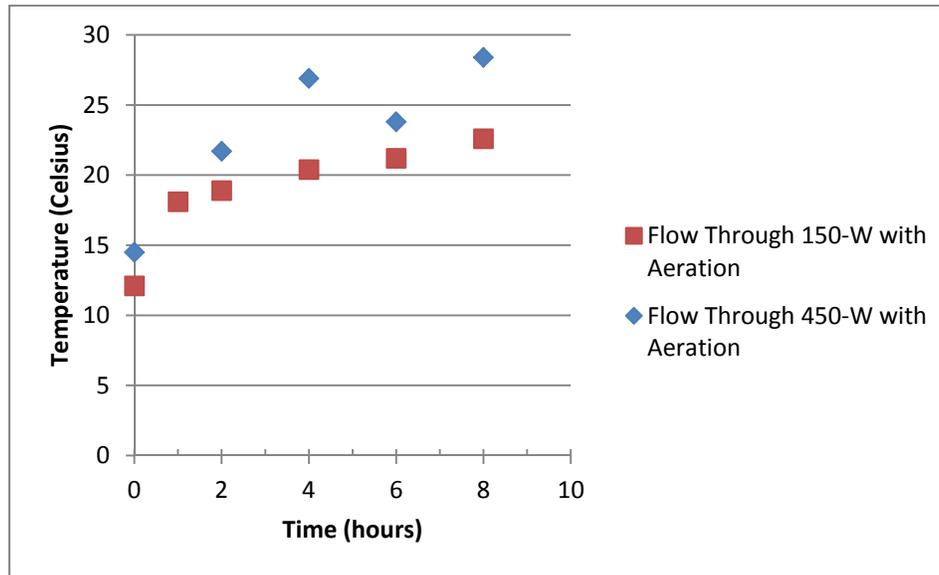
**Table 15. Flow through reactor experiment with aeration using 150-W light source ( $\text{TiO}_2 = 10\text{-g/L}$ ).**

Sample	Temp. °C	COD mg/L	COD % remaining	$\text{NH}_3$ mg/L	$\text{NH}_3$ % remaining	pH	DO mg/L	Alkalinity mg/L as $\text{CaCO}_3$	Alkalinity % remaining
0	12.1	263	100.0	193	100.0	8.70	9.14	900	100.0
1	18.1	263	100.0	135	69.7	8.62	9.01	680	75.6
2	18.9	254	96.6	134	69.4	8.63	8.81	680	75.6
4	20.4	246	93.5	129	66.6	8.71	8.90	680	75.6
6	21.2	255	97.0	126	65.3	8.81	8.81	630	70.0
8	22.6	243	92.4	119	61.7	9.00	8.31	620	68.9

**Table 16. Flow through reactor experiment with aeration using 450-W light source ( $\text{TiO}_2 = 10\text{-g/L}$ ).**

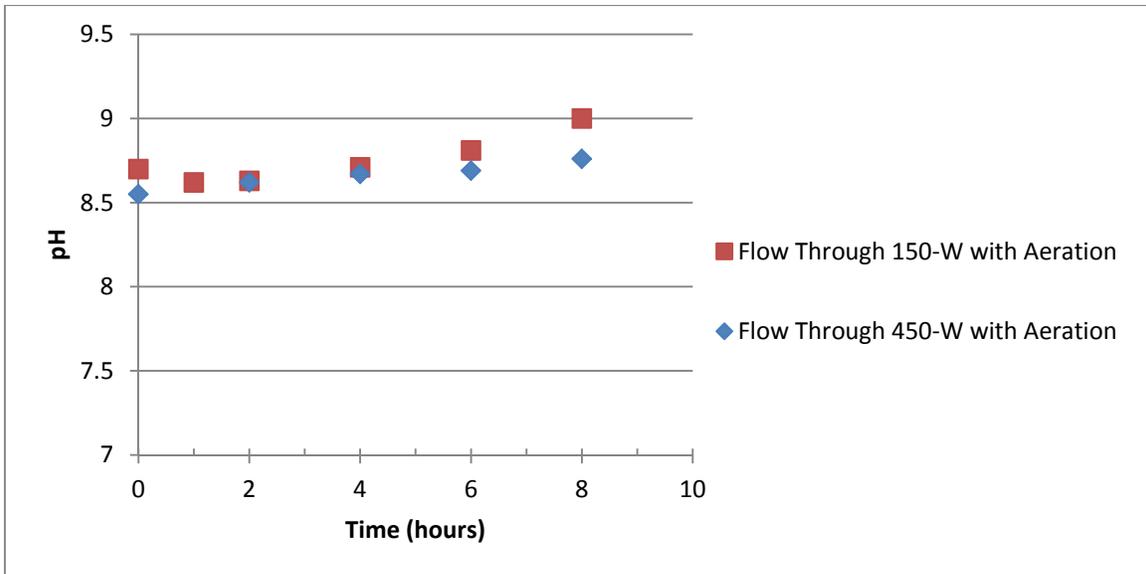
Sample	Temp. °C	COD mg/L	COD % remaining	$\text{NH}_3$ mg/L	$\text{NH}_3$ % remaining	pH	DO mg/L	Alkalinity mg/L as $\text{CaCO}_3$	Alkalinity % remaining
0	14.5	240	100.0	100	100.0	8.55	9.34	700	100.0
2	21.7	220	91.7	95	95	8.62	8.57	700	100.0
4	26.9	217	90.4	84	84	8.67	8.13	490	70.0
6	23.8	210	87.5	77	77	8.69	8.53	430	61.4
8	28.4	189	78.8	75	75	8.76	7.85	700	100.0

The flow through reactor experiment with aeration temperature trend (Figure 27) shows the ability to keep the temperature nearly constant for the duration of the test. Both lamps ran slightly hotter due to the lower fluid levels from reusing the same leachate over again. With the 450-W lamp reaching 28.4°C and the 150-W lamp reaching 22.6°C, the reaction rate would be expected to double in comparison by using the Arrhenius relationship.



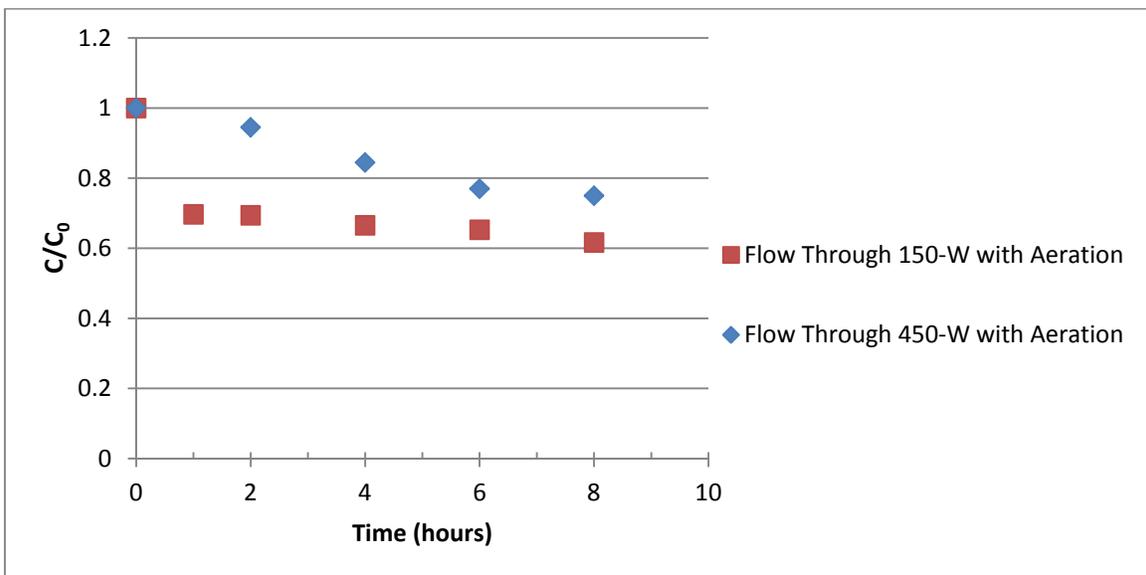
**Figure 32. Flow Through Reactor with Aeration Temperature Comparison**

The flow through reactor experiment with aeration pH trend (Figure 33) started with the same pH as the previous non-aerated flow through reactor experiment #2. The pH remained in the range of 8.5 to 9.0 for the 450-W lamp, and for the 150-W lamp, the pH also remained in the range of 8.5 to 9.0. Both lamps showed no appreciable change in pH during the test.



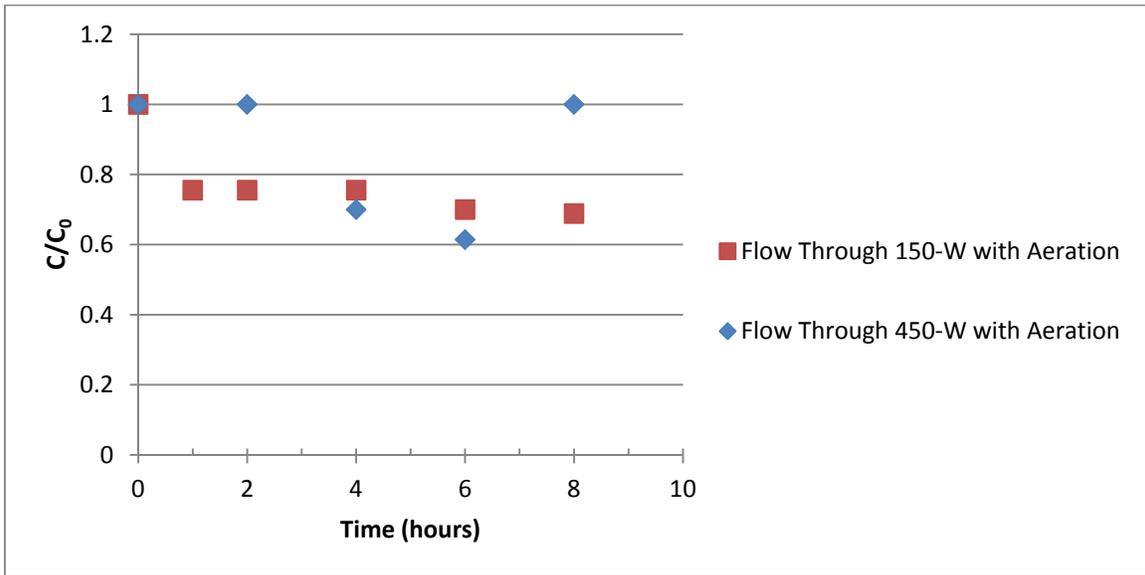
**Figure 33. Flow Through Reactor with Aeration pH Comparison.**

The flow through reactor experiment with aeration ammonia trend (Figure 34) showed a 25.0% reduction in ammonia for the 450-W lamp, compared to the 45.3% reduction from the previous 450-W non-aerated flow through reactor experiment #2. The 150-W lamp showed a reduction of 38.3% compared with the 150-W non-aerated flow through reactor experiment #2, which showed only 26.8% removal. The limited removal for the 450-W lamp may be due to the low starting concentration when compared to the 150-W lamp. The 450-W lamp run started with an ammonia concentration of 100 mg/L compared to the 150-W lamp, which started at 193 mg/L, nearly double.



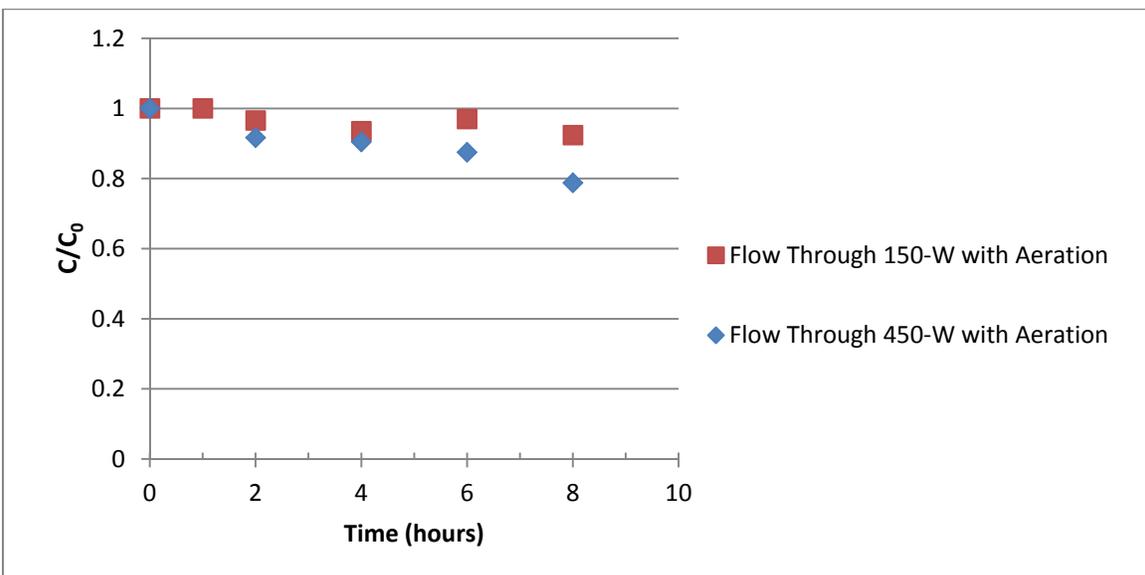
**Figure 34. Flow Through Reactor with Aeration Ammonia Comparison.**

The flow through reactor experiment with aeration alkalinity trend (Figure 30) shows that the 450-W light removed up to 40% of the alkalinity compared to the previous non-aerated flow through test, which removed 45.3% of the alkalinity. The 150-W light removed 31.1% compared to the non-aerated flow through test, which only removed 16.8%



**Figure 35. Flow Through Reactor with Aeration Alkalinity Comparison.**

The flow through reactor experiment with aeration COD trend (**Error! Reference source not found.**) showed only a small removal from both the 450-W and the 150-W light. The 450-W light removed 21.2 % compared with the non-aerated flow through test removal of 7.8%. The 150-W light removed 27.6%, compared to the previous non-aerated experiment removal of 25%.



**Figure 36. Flow Through Reactor with Aeration COD Comparison.**

This was the first run of the flow through reactor with aeration directly in the reaction zone. The best removal achieved was 39% of ammonia by the 150-W lamp. The removal of COD is not below the 100-mg/L goal nor is the alkalinity to the 20-mg/L goal. Nevertheless, improvements were made in removal percentages compared to the previous reactor configuration.

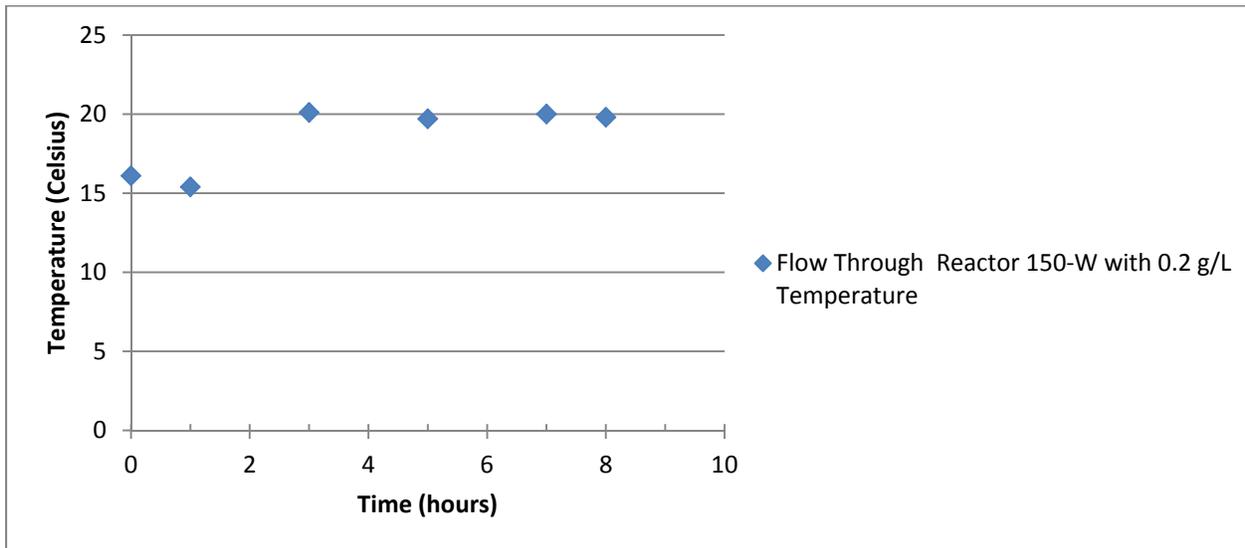
#### **3.4 FLOW THROUGH WITH AERATION EXPERIMENT #4 WITH 0.2-g/L TiO<sub>2</sub> USING THE 150-W LAMP**

This experiment was run with the 150-W lamp and with the addition of an anti-foaming agent, described in the methodology section. The test was run for 8-hours with samples taken at 2-hour intervals. The TiO<sub>2</sub> concentration was reduced from 10 g/L to 0.2-g/L. The maximum COD removal for the trial was 13%. The maximum ammonia removal was 11%. The maximum alkalinity removal was 30%. A summary of results is presented in Table 17.

**Table 17. Flow through reactor experiment with aeration using 150-W light source (TiO<sub>2</sub> = 0.2-g/L).**

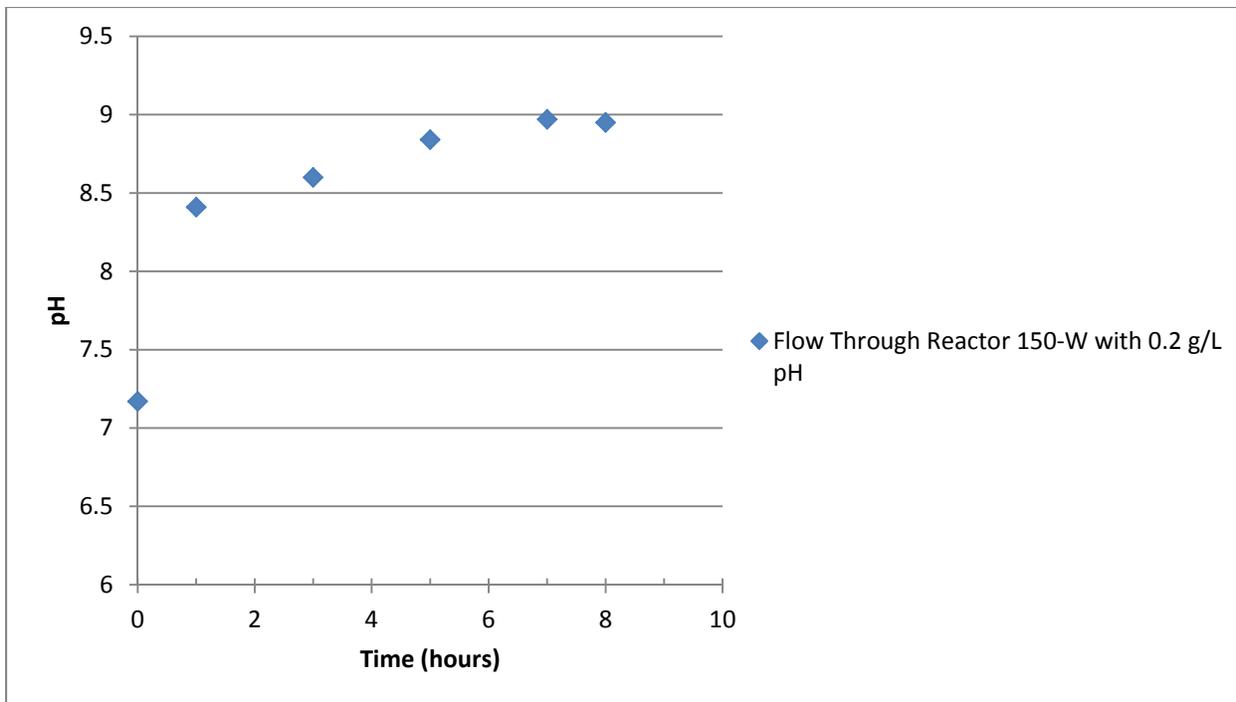
<b>Sample</b>	<b>Temp. °C</b>	<b>COD mg/L</b>	<b>COD % remaining</b>	<b>NH<sub>3</sub> mg/L</b>	<b>NH<sub>3</sub> % remaining</b>	<b>pH</b>	<b>DO mg/L</b>	<b>Alkalinity mg/L as CaCO<sub>3</sub></b>	<b>Alkalinity % remaining</b>
Raw	19.1	294	--	194	100.0	7.17	6.48	1310	100.0
Filtered	17.9	312	--	192	99.2	7.26	7.41	1290	98.5
Anti-Foam	16.1	441	100.0	190	97.9	7.93	10.06	1240	94.7
1	15.4	427	96.8	181	93.5	8.41	9.90	1100	84.0
3	20.1	408	92.5	188	96.9	8.60	9.01	960	73.3
5	19.7	395	89.6	177	91.5	8.84	9.09	920	70.2
7	20.0	388	88.0	184	95.1	8.97	8.97	910	69.5
8	19.8	388	88.0	174	89.7	8.95	9.08	920	70.2

The flow through reactor experiment with aeration using 0.2 g/L TiO<sub>2</sub> temperature trend (Figure 37) show the ability to maintain the temperature near constant for the during of the test. But the ability to keep the temperature constant brings in question the effect of the temperature on the removal process. A higher temperature would be expected to cause hardness to fall out of solution, reducing alkalinity in the process. Higher temperatures would also affect the amount of ammonia that the leachate can hold in bulk solution. This effect could potentially be responsible for more removal than the actual TiO<sub>2</sub> and UV treatment accounts for.



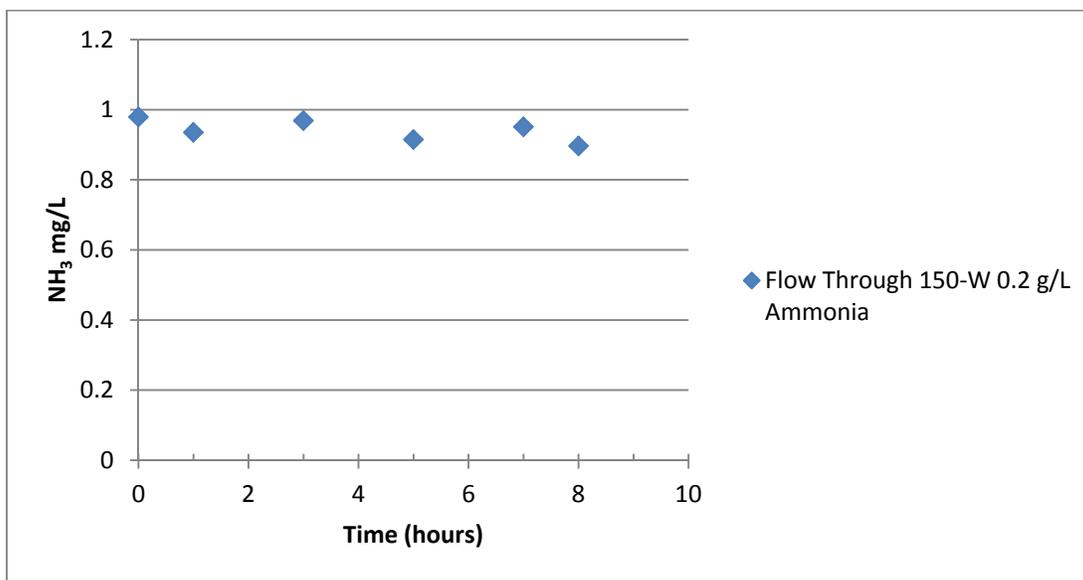
**Figure 37. Flow Through Reactor with 150-W Lamp and Aeration using TiO<sub>2</sub> = 0.2 g/L Temperature Trend.**

The flow through reactor experiment with aeration using 0.2 g/L TiO<sub>2</sub> pH trend (Figure 38) started with a pH of 7.93. The pH increased to 8.97 during the test. The last two readings were almost the same with 8.97 and 8.95. This shows that during the test, the pH does increase, but if the organics were reacting with the photocatalytic process, the pH should be falling, because of the production of CO<sub>2</sub>. The alkalinity of the leachate was reduced, making it easier to change pH of the leachate. It was hypothesized that the generation of CO<sub>2</sub> from the reduction of organic compound would reduce the pH. However, the pH increase, bringing into question if CO<sub>2</sub> corrections.



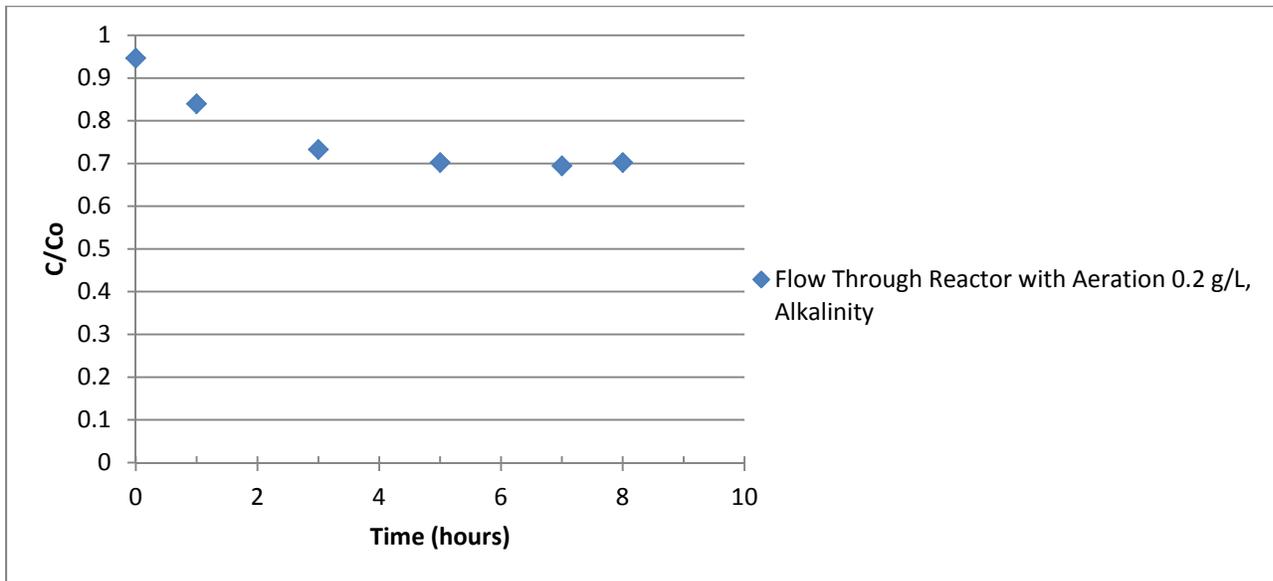
**Figure 38. Flow Through Reactor with 150-W Lamp and Aeration using TiO<sub>2</sub> = 0.2 g/L pH Trend.**

The flow through reactor experiment with aeration using 0.2 g/L TiO<sub>2</sub> ammonia trend (Figure 39) showed a 10.4% reduction in ammonia, compared to the 40.0% reduction from the previous 150-W flow through test using 10 g/L of photocatalyst. The exact reason for this reduction in efficiency is unknown, but the initial hypothesis is that the TiO<sub>2</sub> dosage does play a role in the effectiveness of the removal.



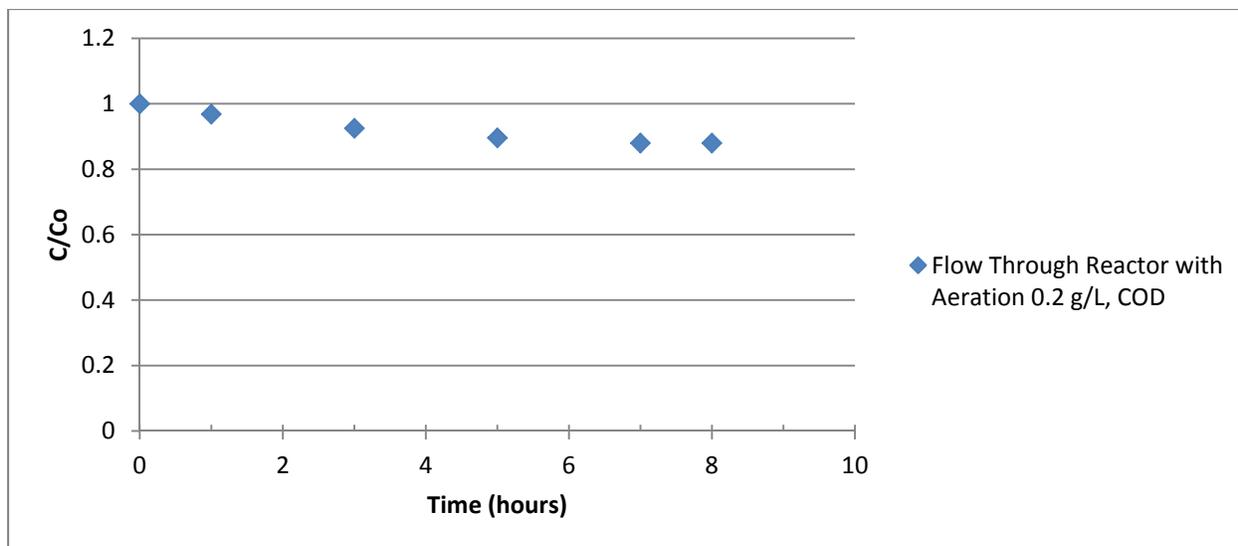
**Figure 39. Flow Through Reactor with 150-W Lamp and Aeration using  $\text{TiO}_2 = 0.2 \text{ g/L}$  Ammonia Trend.**

The flow through reactor experiment with aeration using  $0.2 \text{ g/L TiO}_2$  alkalinity trend (Figure 40) shows that the 150-W light removed 30.0% of the alkalinity compared to the previous flow through test which removed only 31.1%, which used  $10 \text{ g/L}$  of photocatalyst. These removal rates are almost identical and show that the  $\text{TiO}_2$  dosage may not be the major contributing factor to the ammonia removal.



**Figure 40. Flow Through Reactor with 150-W Lamp and Aeration using  $\text{TiO}_2 = 0.2 \text{ g/L}$  Alkalinity Trend.**

The flow through reactor experiment with aeration using  $0.2 \text{ g/L TiO}_2$  COD trend (Figure 41) showed 11.3% removal. Compare this to the previous experiment's 7.7% removal with  $10 \text{ g/L}$  of photocatalyst, and the result is only slightly better. What is not seen in Figure 41 is that the anti-foaming agent was added to the reactor to make controlling the process easier. This anti-foaming agent raised the starting COD from  $294 \text{ mg/L}$  for raw leachate to  $441 \text{ mg/L}$  for the leachate treated with anti-foam. The results of the test show that the final COD removal was still above the raw leachate by 32% because of the addition of the anti-foaming agent.

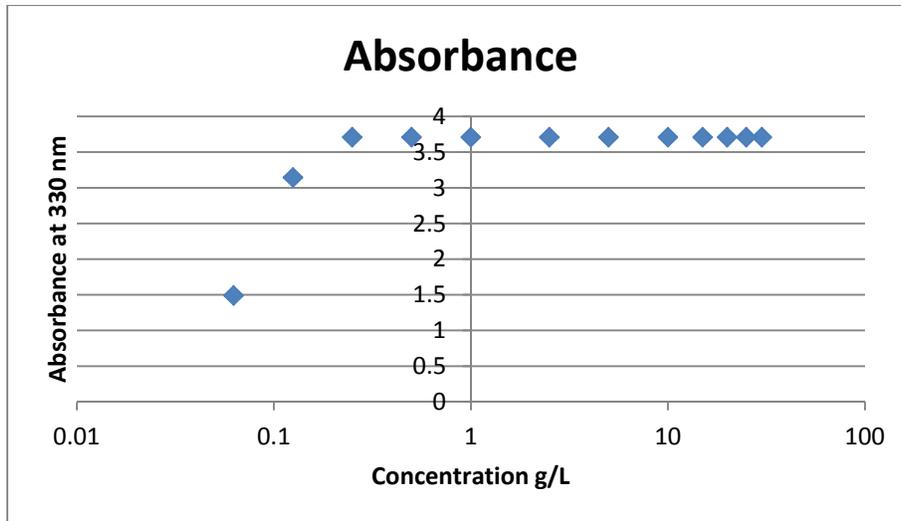


**Figure 41. Flow Through Reactor with 150-W Lamp with Aeration using  $\text{TiO}_2 = 0.2 \text{ g/L}$  COD Trend.**

This was the first run that used anti-foaming agents, and it did make operating the reactor easier. However, it did increase the COD demand by about 130 mg/L after it was added. This was concerning because this is one of the parameters that is the focus to eliminate, after the 8-hour run, the COD levels had not returned to their original levels in the raw leachate. The best removal was 30% of alkalinity. This test showed improvements in the treatment process.

### 3.5 $\text{TiO}_2$ DOSING TEST

During testing, a question about how much light penetration is achieved came up. Since the flow through reactor light waves must travel 2 inches to reach the full bulk solution, the light absorbance of the  $\text{TiO}_2$  at various concentrations was measured. In the falling film reactor, light penetration is not an issue because the solution falls in a thin sheet. Twelve cuvettes were prepared with different dilutions to scan at 330-nm (refer to Figure 15), which was the maximum observed absorbance wavelength. The results are plotted in Figure 42.



**Figure 42. Absorbance curve of TiO<sub>2</sub> dosing.**

Figure 42 shows that the maximum absorbance within the range of the spectrometer, without overdosing TiO<sub>2</sub> such that UV light is blocked, is between 125 and 250 mg/L. This idealized amount of TiO<sub>2</sub> is much lower than what was used in the first six runs. This does not mean that no UV light passed through the cuvette at the higher dilutions but that the TiO<sub>2</sub> cannot absorb any more UV light at the intensity of the spectrophotometer light source, which is different from the lamps used for pilot testing.

### 3.6 CRITICAL ORIFICE ADVANCED OXIDATION PROCESS TESTING

An experiment was run with the critical orifice advanced oxidation unit. The test lasted for 20-minutes, and samples were taken at 5-minute intervals. The test showed impressive results as seen in Table 18.

**Table 18. Alternative AOP test results.**

Sample	DO (mg/L)	TDS (g/L)	pH	Total Alkalinity (mg/L as CaCO <sub>3</sub> )	Conductivity (mS/cm)	COD (mg/L)
Treated Tap Water	5.25	0.437	7.99	230	0.673	281
Leachate Sample 1	5.63	0.376	8.02	230	0.579	103
Leachate Sample 2	5.97	0.587	7.63	210	0.903	91
Leachate Sample 3	6.12	0.669	7.71	330	1.029	79
Leachate Sample 4	5.53	0.677	7.75	360	1.026	83

The device showed COD removal within ten minutes of treatment. The COD levels were below the target goal level of 120-mg/L in 10-minutes. However, the alkalinity rose from 230-mg/L as CaCO<sub>3</sub> to 360-mg/L as CaCO<sub>3</sub>, as the COD fell. The magnetic chambers did not remove any calcium; however, there is some suggestion that the crystal structure of the calcium bicarbonate may be re-arranged into a longer chain, and the structure of the calcium may change from calcite to aragonite (Coe and Cass 2000). The aragonite calcium has a long linear shape increasing

surface area and may make it possible for the TiO<sub>2</sub> to process the calcium more effectively because of the surface area.

### 3.7 CRYSTAL VIOLET TEST

This test was performed to confirm that the TiO<sub>2</sub> process was generating hydroxyl radicals. Crystal violet exhibits a single proton replacement reaction, so it should show that hydroxides are being released and decomposing the crystal violet. A photograph of the petri dish taken before the test can be seen in Figure 43.



**Figure 43. Petri dish with TiO<sub>2</sub> and crystal violet.**

It was expected that the color would disappear in 20 minutes, but after the first trial and checking on it after five minutes, it was discovered that the reaction had already taken place. A second trial was conducted after one minute, and surprisingly the reaction was complete in less than one minute. It was determined that hydroxides were being produced in large quantities, and the reactions for TiO<sub>2</sub> were working.

### 3.8 IMPROVING COD REMOVAL

With the focus for the first part of the research being on methods for improving the removal of COD, additives that could potentially increase the removal efficiency were tested. It was hypothesized that the complete chain of the TiO<sub>2</sub> reaction may not be taking place. That not enough metals were in the leachate to collect electrons, thus preventing the formation of hydroxyl radicals by the TiO<sub>2</sub> not wanting to obtain electrons. To test this hypothesis of preventing short circuiting, certain key metals were added. The results are summarized in Table 19.

**Table 19. Summary of results from using metal additives to stimulate COD removal.**

Sample	COD Final	C/C <sub>o</sub>
Leachate + TiO <sub>2</sub>	258	0.83
Leachate + Zinc + TiO <sub>2</sub> solution	262	0.84
Leachate + UV	276	0.88
Leachate + Zinc + TiO <sub>2</sub> coated aluminum+ TiO <sub>2</sub> solution	276	0.88
Leachate + TiO <sub>2</sub> coated aluminum + TiO <sub>2</sub> solution	292	0.94
Raw Leachate	312	1.00
Leachate + TiO <sub>2</sub> coated aluminum	312	1.00
Leachate + Steel Wool + TiO <sub>2</sub> solution	315	1.01
Leachate + Steel Wool	317	1.02

Leachate + Steel Wool + TiO <sub>2</sub> coated aluminum + TiO <sub>2</sub> solution	319	1.02
Leachate + Zinc	340	1.09

From the results, it is clear that the TiO<sub>2</sub> alone is the best, and that there is no short circuiting happening in the process with fresh leachate. Both the zinc and the steel wool caused the COD to rise, but titanium with zinc achieves a similar COD removal of 84% to the titanium and UV.

### **3.9 INTENSITY OF UV LIGHT**

Calculations for energy that the titanium dioxide was exposed to per liter were made for both the falling film and the flow through reactor. The energy available for titanium dioxide to absorb in conjunction with exposure time should be directly related to efficiency. The results can be seen in Table 20 and able 21.

**Table 20: Falling Film UV Light Intensity.**

Light	Arch Length (cm)	Circumference (cm)	Area (cm <sup>2</sup> )	Measured UV (mW/cm <sup>2</sup> )	Total Power (W)	Test Size (L)	Falling time (Seconds)	Expsoure time (Seconds)	Recirculated Rate (L/hr)	Number of Pass per liter per hour	Expsoure Time per hour (Seconds)	Energy per Hour per liter (W/L)	Length of test (hours)	Total Watt Hours/ Liter (Whr/L)
50W UV A&C	27.94	15.70	438.66	56.00	24.56	9.30	0.41	0.14	320.00	34.41	4.75	0.03	8.00	0.26
450W UV C	27.94	15.70	438.66	0.06	0.03	9.30	0.41	0.14	320.00	34.41	4.75	0.00	8.00	0.00
50W UV A&B	79.30	15.70	1245.01	0.50	0.62	10.00	0.41	0.39	320.00	32.00	12.54	0.00	8.00	0.02
150W UV C	79.30	15.70	1245.01	7.21	8.98	10.00	0.41	0.39	320.00	32.00	12.54	0.03	8.00	0.25

**Table 21: Flow Through UV Light Intensity.**

Light	Arch Length (cm)	Circumference (cm)	Area (cm <sup>2</sup> )	Measured UV (mW/cm <sup>2</sup> )	Total Power (W)	test size (L)	Retension Time ( hr)	Percent Expsoure per Retentions Time	Times recirculated per hour per liter (1/hr)	Expsoure per Liter (1/L)	Watts per Liter (W/L)	Test (hr)	Total Watt Hours/ Liter (Whr/L)
450W UV A&B	27.94	15.70	438.66	56.00	24.56	8.60	0.03	0.34	24.42	0.27	6.63	8.00	53.08
450W UV C	27.94	15.70	438.66	0.06	0.03	8.60	0.03	0.34	24.42	0.27	0.01	8.00	0.06
150W UV A&B	79.30	15.70	1245.01	0.50	0.62	9.10	0.03	0.96	23.08	0.72	0.45	8.00	3.61
150W UV C	79.30	15.70	1245.01	7.21	8.98	9.10	0.03	0.96	23.08	0.72	6.50	8.00	52.02

The results show that the leachate in the flow through reactor receives 200 times more light energy per liter in an eight hour test when compared to the falling film reactor. However, this did not translate to greater COD removal with 25% increase in the 150-W and a 9% decrease in the 450-W. The results were similar for ammonia removal with a 10% increase in 150-W and 24% increase for the 450-W.

### **3.10 CATALYST RECOVERY**

By visual inspection, the flocculation test took 9 minutes for all of the  $\text{TiO}_2$  to settle. This is a setting rate of 2.5-mm per minute. Decanting the water from the top caused a small current, and the  $\text{TiO}_2$  needed additional time to re-settle before further decanting could take place. The test with the 1-micron filter bag resulted in the  $\text{TiO}_2$  flowing right through it. This was expected because the  $\text{TiO}_2$  has an average size of 20-nanometers, which is 98% smaller than 1-micron. This test was not successful, but filtration with tighter pore sizes could be used as a finishing step to the treatment subject to inorganic fouling from calcium carbonate scale.

## 4. CONCLUSION

### 4.1 SUMMARY

If leachate were directly discharged into the environment it could have catastrophic effects on both the water and soil. The largest concern is the COD, Ammonia, heavy metals, and TDS in the leachate. The research has shown that to this point the  $\text{TiO}_2$  advanced oxidation process cannot remove enough of the contaminants in a timely manner. However, improvements have been shown in the treatment process and with further research, and a possibly combining a few techniques the complete treatment for discharge can be achieved.

The critical orifice system showed excellent COD removal in 10 minutes. This technique could be combined with  $\text{TiO}_2$  for heavy metal removal and further COD removal. A possible addition of lime softening could remove the alkalinity making leachate treated to acceptable discharge levels. This is the focus of the research that will take place in year 2 of the study.

### 4.2 RECOMMENDATIONS

For the objective of improving removal of constituents from leachate, a complete process will be tested, since removal using only a single process has yet to be successful in the pilot scale. The addition of a 1-micron filter back at the exit of the reactor may filter out any of the crystalized by-products of the reaction not allowing them to be reprocessed and reduce the effectiveness of the removal process. Furthermore, the addition of zinc dust to the filter may change the electric charge of the leachate, helping with hydroxyl radical production.

The data so far indicate that the photocatalytic oxidation process alone may not be able to reach the treatment targets specified in the objectives. One option for improving the removal efficiency of COD is to pretreat with ozone. This should increase the oxidation of organic materials on the leachate. Another option that has shown some success in previous work (Meeroff 2010) will be to pretreat with lime. This should remove the hardness and change the alkalinity, with lower alkalinity it is thought that the reduction of more organic materials will be possible, because this should theoretically reduce the hydroxide radical scavenging by carbonates and calcium before the radicals can interact with organics or other metals. A third process option will be to pretreat with the critical orifice unit; this should create micro-bubbles that could flow directly into the photocatalytic reaction chamber, increasing the surface area for the reaction with the  $\text{TiO}_2$ . This should also make the focus of the  $\text{TiO}_2$  reaction the removal of ammonia and alkalinity.

Finally another option is pretreatment with a combined photoelectrocatalytic approach, where a current in addition to the  $\text{TiO}_2$  will reduce the leachate. From (An, et al. 2001) the use of electrodes increase the reduction of methylene blue by 21.8% and COD by 14.1% in thirty minutes. This process is thought to reduce the recombination of the electron for the  $\text{TiO}_2$ . If these same results for the reduction of COD can be expected for the entire 8-hour process could cause a remarkable improvement in the removal process.

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